

100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 react with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 ?



Answer: $1 \text{ mole of } \text{H}_2\text{CO}_3$ reacts with 1 mole of HCO_3^- to form 1 mole of H_2O and 1 mole of CO_2 .

101) The only intracellular buffer in red blood cells is hemoglobin.

102) Describe how hemoglobin buffers acid.

Hemoglobin buffers acid by binding H^+ ions. The binding of H^+ ions to hemoglobin shifts the equilibrium to the right, releasing O_2 from hemoglobin. This is known as the Bohr effect.

103) Describe how hemoglobin buffers base.

Hemoglobin buffers base by releasing H^+ ions. The release of H^+ ions from hemoglobin shifts the equilibrium to the left, binding O_2 to hemoglobin. This is known as the Bohr effect.

104) Hemoglobin is a conjugate acid-base buffer system.

105) Write the chemical formula for hemoglobin.

106) What is the conjugate base of hemoglobin?



Answer To F321 Chemistry May Series

R Barnett



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