Question 1

- 1(a)(iii) Difference: With powdered CaCO₃ the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid. Similarity: The total volume of CO2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO₃ was used in both experiments.
- $1(a)(iv) CaCO_{3(s)} + 2HNO_{3(aq)} \rightarrow Ca(NO_3)_{2(aq)} + H_2O_{(i)} + CO_{2(g)}$
- 1(a)(v) Moles of CaCO₃ used = (1.8/100) mol. = 0.018 mol.
 From the balanced equation for the reaction, 1 mole CaCO₃ produces 1 mol CO₂.
 Hence volume of gas produced at R.T.P. = (0.018 × 24000) cm³ = 432 cm³.
- 1(a)(vi) There were impurities in the CaCO, used.
- 1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.
- 1(b)(i) Fe3+ present.
- 1(b)(ii) Magnesium ribbon decreases in size/dissolves. Solution becomes pale green.
- 1(b)(iii) $Fe^{2+}_{(ac)} + 2OH^{-}_{(ac)} \rightarrow Fe(OH)_{2(c)}$
- 1(b)(iv) SO₄2-(ac) SO₃2-(ac) and CO₃2-(ac) absent.
- 1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

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