

BASIC CHEMISTRY – Part 2

Name: Answer Key

Date: _____

Pd: _____

PART 1: MIXTURES:

Fill in the blanks from the information provided on the presentation.

- A mixture is two or more substances blended together (not bonded).
- Each substance in a mixture keeps all of its chemical and physical properties.
- Mixtures, unlike compounds, can be separated into their individual substances by physical means.
- Example: salt and water can be separated by evaporation.
- A solution is a mixture of two or more substances, one dissolved in another.
- Two parts of a solution are the solute and the solvent.
- Solute – substance being dissolved
Examples: sugar, salt
- Solvent – substance doing the dissolving
Example: water
- A solution becomes saturated when the solvent holds all the solute that it can. It can no longer dissolve solute.
- Suspensions are a mixture of particles scattered throughout another. The particles do not dissolve.
Examples: clouds, dust, fog, smoke, mud, blood

Reflect:

1. Give another example of solutions including the solute and solvent.
Vinegar and Calcium Carbonate, etc...
2. Can solutions be separated by chemical means?
Perhaps, but will still involve physical separation as well

PART 2: COMPOUNDS

- Compounds are two or more elements chemically joined by bonds.
- They are represented by chemical formulas.
- Properties of compounds are usually different than the atoms.
Example: Hydrogen(gas) + Oxygen(gas) = Water (liquid)
Example: Iron + Oxygen = Rust
- Compounds can be separated into their elements only if the chemical bonds are broken. This can be done by chemical means (chemical reaction)
- A molecule is the smallest part of a compound—like water—that has all the properties of the compound.

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