

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
Conducts electricity Contains hydrogen ion Proton donor Turns litmus paper red React with most metals Releases hydronium Ex: lemon Vinegar Battery acid (sulfuric acid)	conducts electricity contains hydroxide ion proton acceptor turns litmus paper blue releases hydroxide ion Ex: Windex bleach soap

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A** produce hydrogen ions (H^+) in solution
- B** have a bitter taste
- B** produce hydroxide ions (OH^-) ions
- A** react with bases to form salts
- A** have a sour taste
- B** have a slippery, soapy feel
- B** react with acids to form salts
- A** react with many metals

3. Write an equation for the dissociation of the following **acids** in water: follow example

- | | |
|---|-------------------------------------|
| a) $HClO_4 \rightarrow H^+ + ClO_4^-$ | d) $H_2S \rightarrow 2H^+ + S^{2-}$ |
| b) $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$ | e) $HCl \rightarrow H^+ + Cl^-$ |
| c) $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$ | f) $HNO_3 \rightarrow H^+ + NO_3^-$ |

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- | | |
|---|---|
| a) $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$ | d) $KOH \rightarrow K^+ + OH^-$ |
| b) $NaOH \rightarrow Na^+ + OH^-$ | e) $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$ |
| c) $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$ | f) $LiOH \rightarrow Li^+ + OH^-$ |

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



1 Describing Acids And Bases Section Review Answers

Terry Des Jardins, George G. Burton



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Table of Contents 1 Describing Acids And Bases Section Review Answers

1. Understanding the eBook 1 Describing Acids And Bases Section Review Answers
 - The Rise of Digital Reading 1 Describing Acids And Bases Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 1 Describing Acids And Bases Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1 Describing Acids And Bases Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1 Describing Acids And Bases Section Review Answers
 - Personalized Recommendations
 - 1 Describing Acids And Bases Section Review Answers User Reviews and Ratings
 - 1 Describing Acids And Bases Section Review Answers and Bestseller Lists
5. Accessing 1 Describing Acids And Bases Section Review Answers Free and Paid eBooks
 - 1 Describing Acids And Bases Section Review Answers Public Domain eBooks
 - 1 Describing Acids And Bases Section Review Answers eBook Subscription Services
 - 1 Describing Acids And Bases Section Review Answers Budget-Friendly Options
6. Navigating 1 Describing Acids And Bases Section Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 1 Describing Acids And Bases Section Review Answers Compatibility with Devices
 - 1 Describing Acids And Bases Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1 Describing Acids And Bases Section Review Answers
 - Highlighting and Note-Taking 1 Describing Acids And Bases Section Review Answers
 - Interactive Elements 1 Describing Acids And Bases Section Review Answers
8. Staying Engaged with 1 Describing Acids And Bases Section Review Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 1 Describing Acids And Bases Section Review Answers
- 9. Balancing eBooks and Physical Books 1 Describing Acids And Bases Section Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1 Describing Acids And Bases Section Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1 Describing Acids And Bases Section Review Answers
 - Setting Reading Goals 1 Describing Acids And Bases Section Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1 Describing Acids And Bases Section Review Answers
 - Fact-Checking eBook Content of 1 Describing Acids And Bases Section Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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