

Name: _____

Your test is on Friday, 12/18/15

Chemical Reactions Study GuideBalancing Chemical Equations

The first step to balancing chemical reaction equations is to be able to accurately count the number of each type of element on both the reactants and products side of the arrow. This is done by multiplying the **coefficient** (big number in front of an element symbol) with the **subscript** (little number found after the element symbol). **Remember:** if no coefficient or subscript is written, then there is an implied, imaginary "1".

Examples: 3Br_2 $\leftarrow$ there are 6 bromine atoms ($3 \times 2 = 6$)

CaO $\leftarrow$ there is 1 calcium and 1 oxygen (in each case $1 \times 1 = 1$). The "1"s are implied.

Practice: How many of each type of atom is represented.

3Cl_2 6 Fe 1 2S_8 16 MgF_2 1 Mg; 2F

I_2 2 4Ni 4 $2 \text{Li}_2\text{O}$ 4 Li; 2 O H_2O_2 2 H; 2O

H_2SO_4 2 H; 1 SO₄ $2 \text{Pb}(\text{NO}_3)_2$ 2 Pb; 4 NO₃

Now that we've learned to count, it's time to put our skills into action by altering the numbers of certain elements to suit our desires. In any case, our desires are to have the same number of each type of element on both the **products** and **reactants** side of the equation. We do this by **changing the coefficients**.

Practice: Balance each of the following chemical reaction equations. Check your answer by counting each element on each side when you think you're done.



1section Review Describing Chemical Reactions Answer Key

Robert Spilman Turner



1section Review Describing Chemical Reactions Answer Key:

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Patai's 1992 Guide to the Chemistry of Functional Groups* Saul Patai, 1992-04-08 Patai's 1992 Guide to the Chemistry of Functional Groups Saul Patai The Hebrew University of Jerusalem Israel Ever since the publication of the first volume of The Chemistry of Functional Groups in 1964 the Patai series has acted as an essential reference source to many researchers By the end of 1991 the series consisted of 50 titles bound in 73 volumes containing nearly 900 chapters written by over 1250 authors The aim of this Guide as was that of the previous edition is to present sufficient material on each of the published chapters to allow the researcher to decide whether these chapters are relevant and useful for his or her purpose and thus worth pursuing in full For those who are familiar with only selected volumes from the series the cross referencing between complementary and related chapters from different volumes will be invaluable The Guide is fully indexed by both subject and author thus making it an essential reference tool for all organic chemists *Watershed and Lake Processes Affecting Surface Water Acid-base Chemistry* Robert Spilman Turner, 1990

Holt Chemistry Ralph Thomas Myers, 2004 *Glencoe Science* McGraw-Hill Staff, 2001-06 **Si Silicon**, 2013-11-11 This volume concludes the coverage of silicon carbide SiC begun in Silicon Supplement Volume B 2 1984 subtitled Silicon Carbide Part I Part I described the physical properties of SiC SiC diodes molecular species in the SiC C gas phase and amorphous silicon carbon alloys The current Part II Silicon Supplement Volume B 3 1986 covers in its initial chapter the Si C phase diagram and in the final chapters the higher order systems of Si and C with additional elements through boron arranged according to the Gmelin system In between some 95% of the volume focusses on SiC beginning with its natural occurrence preparation and formation and purification continuing with its chemical analysis manufacture of special ized forms electrochemistry and chemical reactions and concluding with descriptions of its myriad applications The final applications section covering electronic devices also describes similar applications of the amorphous Si C alloys The successive chapters in this volume are often closely interrelated since it is often necessary to synthesize SiC directly in a form in which it will be applied SiC cannot be melted and cast nor rolled nor drawn nor is it easily electroplated or sintered or purified Silicon carbide first became known to man when E G Acheson in 1891 used an electric current to heat a mixture of clay and carbon to extremely high temperatures **Harcourt Science: Physical science, [grade] 3, units E and F, teacher's ed**, 2000 **First Australasian Conference on Heat and Mass Transfer, Monash University, May 23, 24, 25, 1973**, 1973 Monthly Technical Review, 1957 *Paint, Oil and Chemical Review ...*, 1912

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