

5-2 Review and Reinforcement Part II

Quantum Theory and the Atom

Complete each statement, using the vocabulary and the characteristics and history of quantum theory.

characteristic

Pauli exclusion principle

quantum-mechanical model

quantum-mechanical

electron density

orbital

principal energy levels

_____ orbital _____ 1. A region in space around the nucleus where electrons in atoms are likely to be.

_____ electron density _____ 2. The density of an electron cloud.

_____ quantum-mechanical _____ 3. Another description of a quantum-mechanical model of an atom.

Pauli exclusion principle _____ 4. States that any two electrons cannot occupy the same atomic orbital and that those _____

quantum-mechanical model occupies space.

_____ quantum-mechanical model _____ 5. A mathematical progression obtained by treating the electron as waves and _____

quantifying its energy.

_____ quantum-mechanical _____ 6. The model of an atom that describes _____

electron spin _____ 7. The introduction of quantum-mechanical model of an atom.

Complete each statement, using the history of the quantum-mechanical model of an atom.

_____ 8. _____ 1. The electron cloud is _____ where the probability of finding an electron is _____

- a. greatest
- b. lowest
- c. highest density
- d. intermediate

_____ 9. _____ 2. The low principal energy level of the hydrogen atom contains only one _____

- a. orbital
- b. set of orbitals
- c. set of orbitals
- d. orbital

20 2 Review And Reinforcement Continued Answers

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