3	This question is about atoms, molecules and nanoparticles.
3 (a)	Different atoms have different numbers of sub-atomic particles.
3 (a) (i)	An oxygen atom can be represented as \$^16O
	Explain why the mass number of this atom is 16.
	You should refer to the numbers of sub-atomic particles in the nucleus of the atom. [2 marks]
3 (a) (ii)	Explain why $^{12}_{\ 6}\mathrm{C}$ and $^{14}_{\ 6}\mathrm{C}$ are isotopes of carbon.
	You should refer to the numbers of sub-atomic particles in the nucleus of each isotope. [3 marks]
3 (b)	Hydrogen atoms and oxygen atoms chemically combine to produce water molecules.
3 (b) (i)	Complete Figure 3 to show the arrangement of the outer shell electrons of the hydrogen and oxygen atoms in a molecule of water.
	Use dots (•) or crosses (×) to represent the electrons. [2 marks]
	Figure 3
	н ° н
3 (b) (ii)	Name the type of bonding in a molecule of water. [1 mark]
3 (b) (iii)	Why does pure water not conduct electricity? [1 mark]
3 (c)	Nanoparticles of cobalt oxide can be used as catalysts in the production of hydrogen from water.
3 (c) (i)	How does the size of a nanoparticle compare with the size of an atom? [1 mark]
3 (c) (ii)	Suggest one reason why 1 g of cobalt oxide nanoparticles is a better catalyst than 1 g of cobalt oxide powder. [1 mark]

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- 1. Understanding the eBook Alternative To Practical Chemistry Past Question 2013 Gsce
 - The Rise of Digital Reading Alternative To Practical Chemistry Past Question 2013 Gsce
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Alternative To Practical Chemistry Past Question 2013 Gsce
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Alternative To Practical Chemistry Past Question 2013 Gsce
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Alternative To Practical Chemistry Past Question 2013 Gsce
 - Personalized Recommendations
 - Alternative To Practical Chemistry Past Question 2013 Gsce User Reviews and Ratings
 - Alternative To Practical Chemistry Past Question 2013 Gsce and Bestseller Lists

- 5. Accessing Alternative To Practical Chemistry Past Question 2013 Gsce Free and Paid eBooks
 - Alternative To Practical Chemistry Past Question 2013 Gsce Public Domain eBooks
 - Alternative To Practical Chemistry Past Question 2013 Gsce eBook Subscription Services
 - Alternative To Practical Chemistry Past Question 2013 Gsce Budget-Friendly Options
- 6. Navigating Alternative To Practical Chemistry Past Question 2013 Gsce eBook Formats
 - o ePub, PDF, MOBI, and More
 - Alternative To Practical Chemistry Past Question 2013 Gsce Compatibility with Devices
 - Alternative To Practical Chemistry Past Question 2013 Gsce Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Alternative To Practical Chemistry Past Question 2013 Gsce
 - Highlighting and Note-Taking Alternative To Practical Chemistry Past Question 2013 Gsce
 - Interactive Elements Alternative To Practical Chemistry Past Question 2013 Gsce
- 8. Staying Engaged with Alternative To Practical Chemistry Past Question 2013 Gsce
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Alternative To Practical Chemistry Past Question 2013 Gsce
- 9. Balancing eBooks and Physical Books Alternative To Practical Chemistry Past Question 2013 Gsce
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Alternative To Practical Chemistry Past Question 2013 Gsce
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Alternative To Practical Chemistry Past Question 2013 Gsce
 - Setting Reading Goals Alternative To Practical Chemistry Past Question 2013 Gsce
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Alternative To Practical Chemistry Past Question 2013 Gsce
 - Fact-Checking eBook Content of Alternative To Practical Chemistry Past Question 2013 Gsce
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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