

Assignment #2 – Compound Names and Formulas

Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg₂O
3. gold (III) chloride AuCl₃
4. thallium (III) bromide TlBr₃
5. bismuth (V) oxide Bi₂O₅
6. terbium (III) oxide Tb₂O₃
7. uranium (VI) oxide UO₃
8. protactinium (IV) bromide PaBr₄
9. cerium (III) oxide Ce₂O₃
10. arsenic (V) sulphide As₂S₅

11. manganese (III) oxide Mn₂O₃
12. vanadium (II) bromide VBr₂
13. niobium(V) oxide Nb₂O₅
14. titanium (III) oxide Ti₂O₃
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co₃P₂
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr₂
20. copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl₄ Tin (IV) Chloride
2. BiBr₅ Bismuth (V) Bromide
3. PoO₂ Polonium (IV) Oxide
4. PbI₂ Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au₂O₃ Gold (III) Oxide
8. FeCl₂ Iron (II) Chloride

9. PdF₄ Palladium (IV) Fluoride
10. Os₂O₃ Osmium (III) Oxide
11. MoBr₂ Molybdenum (II) Bromide
12. VCl₅ Vanadium (V) Chloride
13. Mn₂O₃ Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np₂O₃ Neptunium (III) Oxide
16. V₂O₅ Vanadium (V) Oxide

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Table of Contents Answer Key To Writing Ionic Formulas Math Skills

1. Understanding the eBook Answer Key To Writing Ionic Formulas Math Skills
 - The Rise of Digital Reading Answer Key To Writing Ionic Formulas Math Skills
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Writing Ionic Formulas Math Skills
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Writing Ionic Formulas Math Skills
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Writing Ionic Formulas Math Skills
 - Personalized Recommendations
 - Answer Key To Writing Ionic Formulas Math Skills User Reviews and Ratings
 - Answer Key To Writing Ionic Formulas Math Skills and Bestseller Lists
5. Accessing Answer Key To Writing Ionic Formulas Math Skills Free and Paid eBooks
 - Answer Key To Writing Ionic Formulas Math Skills Public Domain eBooks
 - Answer Key To Writing Ionic Formulas Math Skills eBook Subscription Services
 - Answer Key To Writing Ionic Formulas Math Skills Budget-Friendly Options
6. Navigating Answer Key To Writing Ionic Formulas Math Skills eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To Writing Ionic Formulas Math Skills Compatibility with Devices
 - Answer Key To Writing Ionic Formulas Math Skills Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Writing Ionic Formulas Math Skills
 - Highlighting and Note-Taking Answer Key To Writing Ionic Formulas Math Skills
 - Interactive Elements Answer Key To Writing Ionic Formulas Math Skills

8. Staying Engaged with Answer Key To Writing Ionic Formulas Math Skills
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Writing Ionic Formulas Math Skills
9. Balancing eBooks and Physical Books Answer Key To Writing Ionic Formulas Math Skills
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Writing Ionic Formulas Math Skills
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Writing Ionic Formulas Math Skills
 - Setting Reading Goals Answer Key To Writing Ionic Formulas Math Skills
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Writing Ionic Formulas Math Skills
 - Fact-Checking eBook Content of Answer Key To Writing Ionic Formulas Math Skills
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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