

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₂.

$$\frac{9 \text{ FeCl}_2}{1 \text{ mole FeCl}_2} \times \frac{1 \text{ mole FeCl}_2}{9 \text{ FeCl}_2} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_2}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₂ produced with 95g of Fe (SO₄)₂

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_2}{\text{mole Fe}(\text{SO}_4)_2} \times \frac{9 \text{ FeCl}_2}{1 \text{ mole FeCl}_2}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_2}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

Answers For Study Guide Mass And Moles

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
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Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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