

# Hydrolysis -

## The Reaction of Ions with Water

### Lab Report

#### Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

#### Hypothesis

|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         |                             |         |
|-------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------|---------|
| Sodium Acetate    | $\text{NaCH}_3\text{COO}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CH}_3\text{COO}^-_{(aq)}$ $\text{CH}_3\text{COO}^-_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{CH}_3\text{COOH}_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                                                                                                                         |                             | Basic   |
| Sodium Chloride   | $\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$                                                                                                                                                                                                                                                                                                                                                                             |                             | Neutral |
| Ammonium Chloride | $\text{NH}_4\text{Cl}_{(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$                                                                                                                                                                                                                                                 |                             | Acidic  |
| Ammonium Sulfate  | $(\text{NH}_4)_2\text{SO}_{4(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                      | $K_a = 5.4 \times 10^{-10}$ | Acidic  |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 8.3 \times 10^{-13}$ |         |
| Aluminum Chloride | $\text{AlCl}_{3(s)} \rightarrow \text{Al}^{3+}_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{Al}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Al}(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$                                                                                                                                                                                                                                                         |                             | Acidic  |
| Calcium Nitrate   | $\text{Ca}(\text{NO}_3)_{2(s)} \rightarrow \text{Ca}^{2+}_{(aq)} + \text{NO}_{3(aq)}$                                                                                                                                                                                                                                                                                                                                                                   |                             | Neutral |
| Iron(III) Sulfate | $\text{Fe}_2(\text{SO}_4)_{3(s)} \rightarrow \text{Fe}^{3+}_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Fe}(\text{OH})^{2+}_{(aq)} + \text{H}^+_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HFe}^{2+}_{(aq)} + \text{OH}^-_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$ | $K_a = 6.0 \times 10^{-3}$  | Acidic  |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 1.7 \times 10^{-12}$ |         |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 8.3 \times 10^{-13}$ |         |
| Sodium Carbonate  | $\text{Na}_2\text{CO}_{3(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CO}_3^{2-}_{(aq)}$ $\text{CO}_3^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HCO}_3^-_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                                                                                                                                                 |                             | Basic   |

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**Judd E. Hollander**



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