

# Kinetics Worksheet 1

3. For the reaction:  $A + 2B \rightarrow 3C + D$   
the following data were obtained,

| run | [A]  | [B]  | rate $M \cdot s^{-1}$ |
|-----|------|------|-----------------------|
| 1   | 0.16 | 0.36 | 2.67                  |
| 2   | 0.16 | 0.09 | 0.67                  |
| 3   | 0.32 | 0.24 | 7.11                  |
| 4   | 0.08 | 0.09 | 0.168                 |
| 5   | 0.24 | 0.12 | 2.0                   |
| 6   | 0.02 | 1.35 | ?                     |

$$R = k[A]^y[B]^x$$

- (A) Write the rate law expression for this reaction.

(B)  $\left(\frac{0.36}{0.09}\right)^x = \frac{2.67 \text{ m s}^{-1}}{0.67 \text{ m s}^{-1}}$  (A)  $\left(\frac{0.16}{0.08}\right)^y = \frac{0.67 \text{ m s}^{-1}}{0.168 \text{ m s}^{-1}}$   
 $(4)^x = 4$   $(2)^y = 4$   
 $x \ln 4 = \ln 4$   $y \ln 2 = \ln 4$   
 $x = 1$   $y = 2$

$$R = k[A]^2[B]$$

- (B) Calculate the value of k, the rate constant (remember units).

$$k = \frac{R}{[A]^2[B]} = \frac{2.67 \text{ m s}^{-1}}{(0.16)^2 (0.36)} \quad k = \text{m}^{-2} \text{s}^{-1}$$

$$k = 290 \text{ m}^{-2} \text{s}^{-1}$$

- (C) Calculate the rate of the reaction when  $[A] = 0.02 \text{ M}$  and  $[B] = 1.35 \text{ M}$ .

$$R = 290 \text{ m}^{-2} \text{s}^{-1} [0.02]^2 [1.35] = 0.16 \frac{\text{m}}{\text{s}}$$

- (D) Assume that the reaction goes to completion. Under the conditions specified in the second run, what would the final [C] be?



$$\text{limit} \Rightarrow 0.09 B \times \frac{1A}{2B} = 0.045 A$$

$$0.045 A \times \frac{3C}{1A} = 0.135 \text{ m C}$$

$$\begin{aligned} .16 A \times \frac{3C}{1A} &= .48 \text{ m} \\ .09 B \times \frac{3C}{2B} &= .135 \text{ m} \end{aligned}$$

or R

# Answers Kinetics Worksheet 1

**Arshad Iqbal**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## Answers Kinetics Worksheet 1:

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## **Table of Contents Answers Kinetics Worksheet 1**

1. Understanding the eBook Answers Kinetics Worksheet 1
  - The Rise of Digital Reading Answers Kinetics Worksheet 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers Kinetics Worksheet 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers Kinetics Worksheet 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Kinetics Worksheet 1
  - Personalized Recommendations
  - Answers Kinetics Worksheet 1 User Reviews and Ratings
  - Answers Kinetics Worksheet 1 and Bestseller Lists
5. Accessing Answers Kinetics Worksheet 1 Free and Paid eBooks

- Answers Kinetics Worksheet 1 Public Domain eBooks
- Answers Kinetics Worksheet 1 eBook Subscription Services
- Answers Kinetics Worksheet 1 Budget-Friendly Options
- 6. Navigating Answers Kinetics Worksheet 1 eBook Formats
  - ePub, PDF, MOBI, and More
  - Answers Kinetics Worksheet 1 Compatibility with Devices
  - Answers Kinetics Worksheet 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers Kinetics Worksheet 1
  - Highlighting and Note-Taking Answers Kinetics Worksheet 1
  - Interactive Elements Answers Kinetics Worksheet 1
- 8. Staying Engaged with Answers Kinetics Worksheet 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers Kinetics Worksheet 1
- 9. Balancing eBooks and Physical Books Answers Kinetics Worksheet 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers Kinetics Worksheet 1
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Kinetics Worksheet 1
  - Setting Reading Goals Answers Kinetics Worksheet 1
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Kinetics Worksheet 1
  - Fact-Checking eBook Content of Answers Kinetics Worksheet 1
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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