

Chemistry Final Exam Review Questions

Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$Q_F = 1.84 Q_C + 12$$

speed of light = 3.00×10^8 m/s

2000

$R = 0.082057 \text{ L-atm / deg-mole}$

Planck's constant = $6.63 \times 10^{-34} \text{ J.s}$

15. *Journal of the American Medical Association*, 2000; 283: 2686-2692.

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does **not** describe gases. Which one is it?
 - Expands infinitely.
 - Rigid shape.
 - Easily compressed.
 - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
 - Its density is 5.43 g/cm^3 .
 - It is a silvery-white metal.
 - It is a liquid at 298 K.
 - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of SO_2 . Which one is a **physical** property?
 - It reacts with oxygen to form SO_3^{2-} .
 - It condenses to a colorless liquid at 14°F and atmospheric pressure.
 - It is a strong reducing agent.
 - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
 - 0.04309
 - 0.0430
 - 0.0431
 - 0.43980
- How many significant figures should there be in the answer to the following math equation?
$$2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$$
 - 3
 - 4
 - 5
 - 6
- How many significant figures should there be in the answer to the following math equation?
$$23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$$
 - 2
 - 3
 - 4
 - 5
- At 25°C , one cm^3 of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
 - 0.82 cm^3
 - 8.2 cm^3
 - 0.164 cm^3
 - 16.4 cm^3
- A cube of lead has a volume of 6.61 cm^3 . Its mass is 75.0 grams. What is its density?
 - 3.85 g/cm^3
 - 7.41 g/cm^3
 - 11.3 g/cm^3
 - 22.6 g/cm^3

Answers For Chemistry Final Exam Review 2014

Richard H. Langley, John Moore



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1. Understanding the eBook Answers For Chemistry Final Exam Review 2014
 - The Rise of Digital Reading Answers For Chemistry Final Exam Review 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Chemistry Final Exam Review 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemistry Final Exam Review 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Chemistry Final Exam Review 2014
 - Personalized Recommendations
 - Answers For Chemistry Final Exam Review 2014 User Reviews and Ratings
 - Answers For Chemistry Final Exam Review 2014 and Bestseller Lists
5. Accessing Answers For Chemistry Final Exam Review 2014 Free and Paid eBooks
 - Answers For Chemistry Final Exam Review 2014 Public Domain eBooks
 - Answers For Chemistry Final Exam Review 2014 eBook Subscription Services
 - Answers For Chemistry Final Exam Review 2014 Budget-Friendly Options

6. Navigating Answers For Chemistry Final Exam Review 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Chemistry Final Exam Review 2014 Compatibility with Devices
 - Answers For Chemistry Final Exam Review 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Chemistry Final Exam Review 2014
 - Highlighting and Note-Taking Answers For Chemistry Final Exam Review 2014
 - Interactive Elements Answers For Chemistry Final Exam Review 2014
8. Staying Engaged with Answers For Chemistry Final Exam Review 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemistry Final Exam Review 2014
9. Balancing eBooks and Physical Books Answers For Chemistry Final Exam Review 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Chemistry Final Exam Review 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Chemistry Final Exam Review 2014
 - Setting Reading Goals Answers For Chemistry Final Exam Review 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Chemistry Final Exam Review 2014
 - Fact-Checking eBook Content of Answers For Chemistry Final Exam Review 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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