

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) =14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) +4(2) =14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) =6 O atoms

Product side

4 Fe atoms

2(3) =6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃:4 mol Fe

Answer Key To Stoichiometry Homework Problems

Yan Bai



Answer Key To Stoichiometry Homework Problems:

Answer Key To Stoichiometry Homework Problems Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Answer Key To Stoichiometry Homework Problems**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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