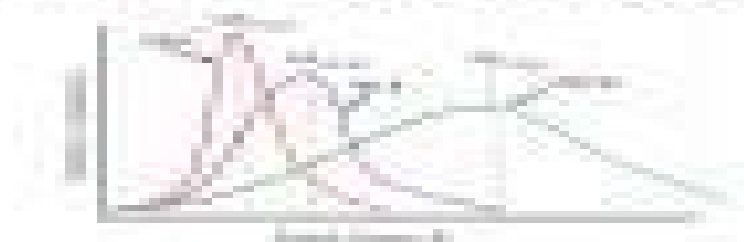


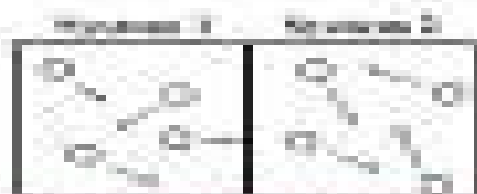
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Change**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving m/s is $\frac{1}{2}mv^2$, the $\frac{1}{2}$ is constant and m is constant, thus the change in kinetic energy can all be temperature increase. The Boltzmann probability curves show that the average velocity is small for the distribution's maximum (peak) to change the average temperature noticeably.

3. Sketch each graph to represent the relative distribution of the particles in the following two gas samples at the same gas:



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are at the same mass, all of the individual particles have the same mass. According to the equation $\text{KE} = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only variable that affects KE is the velocity in System 1 and System 2. Since particles are moving at slower velocities in System 1, the average KE is lower.
- b. Are there different total kinetic energy for System 1? Justify your answer in terms of molecular collisions.
 All gas is transferred from System 1 to System 2. The particles in System 1 are at the same mass and have the same temperature. The gas particles are at the same mass and have the same temperature, so the average KE is the same. The particles in System 1 are at the same mass and have the same temperature, so the average KE is the same. The particles in System 1 are at the same mass and have the same temperature, so the average KE is the same.

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