

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটীক্ষার কপূর, এখানে সেরে পটীক্ষা, আরও কলেজের শিক্ষার্থীর কলেজসমূহের পটীক্ষা এবং মতামত প্রদানের ক্ষমতা
পূর্ণাঙ্গ উত্তর অধ্যয়নিক সেবা প্রদান। এগুলোর অনুশীলন করলে তুমি এ অধ্যয়ন থেকে সেরে পটীক্ষা পূর্ণাঙ্গ কলেজসমূহের
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পানিদ্রুত লবণের প্রাকৃত্য, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রের বিন্দুমান এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রক্রিয়াকৃত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কেন্দ্র পানিদ্রুত লবণের প্রাকৃত্য তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃত্য বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কেন্দ্র পানি দ্রুত লবণ নিষ্কাশিত হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপমাত্রা বৃদ্ধি 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃত্য হ্রাস পড়ে।

২

আমরা জানি,
 $w = MV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, দ্রবীভূত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

১০ পাত্র-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$



সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$
1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর তরুণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর নমুনা তৈরি করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোলার— ব্যাখ্যা করো।
- উদীপকের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন শূন্য ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপূর্ণ তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

উদীপকে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL দ্রবণে HCl এর ভর 2g

$\therefore 1000 \text{ mL} \text{ " " " " } \left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Table of Contents Answer Of Chemistry Mcq Hsc 2015

1. Understanding the eBook Answer Of Chemistry Mcq Hsc 2015
 - The Rise of Digital Reading Answer Of Chemistry Mcq Hsc 2015
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Of Chemistry Mcq Hsc 2015
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Of Chemistry Mcq Hsc 2015
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Of Chemistry Mcq Hsc 2015

- Personalized Recommendations
- Answer Of Chemistry Mcq Hsc 2015 User Reviews and Ratings
- Answer Of Chemistry Mcq Hsc 2015 and Bestseller Lists
- 5. Accessing Answer Of Chemistry Mcq Hsc 2015 Free and Paid eBooks
 - Answer Of Chemistry Mcq Hsc 2015 Public Domain eBooks
 - Answer Of Chemistry Mcq Hsc 2015 eBook Subscription Services
 - Answer Of Chemistry Mcq Hsc 2015 Budget-Friendly Options
- 6. Navigating Answer Of Chemistry Mcq Hsc 2015 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Of Chemistry Mcq Hsc 2015 Compatibility with Devices
 - Answer Of Chemistry Mcq Hsc 2015 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Of Chemistry Mcq Hsc 2015
 - Highlighting and Note-Taking Answer Of Chemistry Mcq Hsc 2015
 - Interactive Elements Answer Of Chemistry Mcq Hsc 2015
- 8. Staying Engaged with Answer Of Chemistry Mcq Hsc 2015
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Of Chemistry Mcq Hsc 2015
- 9. Balancing eBooks and Physical Books Answer Of Chemistry Mcq Hsc 2015
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Of Chemistry Mcq Hsc 2015
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Of Chemistry Mcq Hsc 2015
 - Setting Reading Goals Answer Of Chemistry Mcq Hsc 2015
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Of Chemistry Mcq Hsc 2015

- Fact-Checking eBook Content of Answer Of Chemistry Mcq Hsc 2015
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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