# "PERIODIC TABLE PUZZLE" ANSWER KEY

- 1. C (Cannot be "I" because "I" is Hydrogen, which is a gas, not a metal.)
- F (Group 2 is the Alkaline Earth Metal Family)
- A ("Inactive" means not reactive, also known as "inert." The Noble Gasses in Group 18 are inert because they do not easily react with other elements.)
- H ("Active," meaning "reactive." Group 17, the Halogens, are the most reactive nonmetals. Letter "H" occupies the space of the element Fluorine, which is the most active nonmetal on the Periodic Table.)
- B ("Semi-metal" is the same as a "metalloid," referring to elements that have some of the physical and chemical properties typical of metals, but other properties that are not characteristic of metals.)
- 6. D ("E" is technically a Transition Metal, not an "Inner" Transition Metal)
- 7. G (Oxidation state, or oxidation number, refers to the charge an atom would have if it became an ion. Elements in Group 16, such as "G," all have 6 valence electrons, meaning they would need 2 more electrons to complete their valence shell. Since electrons have negative charges, gaining 2 more would give the ion a charge of -2.)
- E (Transition Metals like "E" do not follow the standard rules. For instance, many have more than one oxidation number. Also, answer can't be "B" because B is a semimetal/metalloid, not a metal.)
- J ("B" is in Group 13, meaning it has 3 valence electrons. It would be easier for it donate 3 negative electrons, leaving it with a charge of +3, than it would for it to gain 5, for a charge of -5.)
- I ("I" is in the place of Hydrogen, which could gain 1 electron to fill the first electron shell, or lose 1 electron to become the common positive ion H\*.)

# **Answer To Chemistry A Study Of Matter Periodic Puzzle**

**Morris Hein, Susan Arena** 

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