Name KEY Pr. Date Write balanced formula equations for these reactions and identify the type of reaction. ▶2KI Potassium + Iodine Synthesis \_\_ 2K + I2 Barium Iodide + Sodium Sulfate \_\_\_\_\_\_ double rep BaIz + Na 2 SO4 -> Ba SO4 + 2NaI Aluminum Hydroxide decomposition 2 AI(OH)3 -> AI2O3 + 3H2O 4. Sodium Chlorate NaCl + Oz decomposition Aluminum Bromide + Chlorine ---single replacement AIBry + Cl2 -> AICI3+ Bry 6. Phosphoric Acid + Magnesium ----H3PO4 + Mg -> Mq3(PO4)2+ H2 7. Iron III Sulfate + Calcium Nitrate ---double rep. Fez(Say)3 +3Ca(NO3)2 -> 2Fe(NO3)2 +3Ca SO4 8. Sodium + Water w Stragte vep ZNa+2H2O -> ZNaOH + HZ 9. Aluminum + Iodine ---synthesis 2 AL +3T2 -> ZALT2 10. Iron III Hydroxide + Hydrobromic Acid double vep. Fe (OH), +3HBr -> FeBr, +3HOH Magnesium Hydroxide + Ammonium Sulfate — double rep Mg(OH). + (NH4) SO4 -> MgSG4 + 2NH4OH 12. Water -12H2+ Dx decomposition 2420 13. Calcium + Hydrochloric Acid ---single repassment ca +2HCI -> Cacis + Hz 14. Lithium + Chlorine -Synthesis 2 Li+Ch -> 2 LiC1

# **1chemical Reactions Practice Problems Key**

SJ Ball

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