

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. reactant
- 4. coefficient
- 5. product

Part A Completion

Write the coefficient for each in the balanced equation, and balance the equation by changing the coefficient for each substance. Write balanced Latin formulas using Roman numerals.

1. A metal compound is reacting according to the formula _____ 2. _____
 The products are water and a metal compound _____ 3. _____ The two
 products have the same number of atoms of each _____ 4. _____
 The two compounds are _____ 5. _____ A chemical reaction between water and
 hydrogen gas is shown. The gas _____ 6. _____ is hydrogen gas.
 Hydrogen gas is the lightest of gases at room temperature. It
 is used in the hydrogen bomb and in the hydrogen fuel cell.
 The chemical equation is _____ 7. _____

8. Hydrogen gas is a colorless, odorless gas. It is used in the hydrogen
 bomb. The chemical equation is _____ 9. _____ for a reaction _____ 10. _____ and for a gas.
 _____ 11. _____ Hydrogen gas is a colorless, odorless gas.
 It is used in the hydrogen bomb and in the hydrogen fuel cell.
 The chemical equation is _____ 12. _____

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____

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