

Chemistry

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1A

FUNDAMENTAL CONCEPTS

5

2011

- Q.1 In mass spectrometer, detector or collector measures the:
 A) Masses of isotopes
 B) Percentages of isotopes
 C) Relative abundances of isotopes
 D) Mass numbers of isotopes
- Q.2 How many ^{35}Cl (chlorine) atoms are in two moles of chlorine?
 A) $2 \times 6.02 \times 10^{23}$ atoms
 B) $35.5 \times 6.02 \times 10^{23}$ atoms
 C) 2×16^{23} atoms
 D) $2 \times 6.02 \times 10^{23}$ atoms

2012

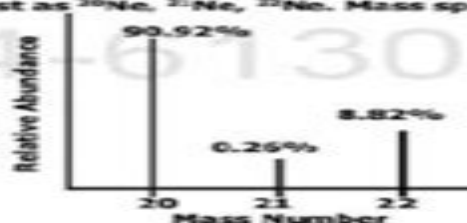
- Q.3 An organic compound has empirical formula $\text{C}_3\text{H}_2\text{O}$, if molar mass of compound is $110.15 \text{ g mol}^{-1}$. The molecular formula of this organic compound is (A, of C=12, H=1.008 and O=16)
 A) $\text{C}_6\text{H}_4\text{O}_2$
 B) $\text{C}_9\text{H}_6\text{O}_3$
 C) $\text{C}_3\text{H}_2\text{O}$
 D) $\text{C}_6\text{H}_4\text{O}_2$
- Q.4 When 8 grams (4 moles) of H_2 react with 2 moles of O_2 , how many moles of water will be formed?
 A) Five
 B) Four
 C) Six
 D) Three

2013

- Q.5 Hydrogen burns in chlorine to produce hydrogen chloride. The ratio of masses of reactants in chemical reaction is:

$$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$$

 A) 1:35.5
 B) 2:35.5
 C) 1:71
 D) 2:70
- Q.6 A sample of Neon is found to exist as ^{20}Ne , ^{21}Ne , ^{22}Ne . Mass spectrum of 'Ne' is as follow:

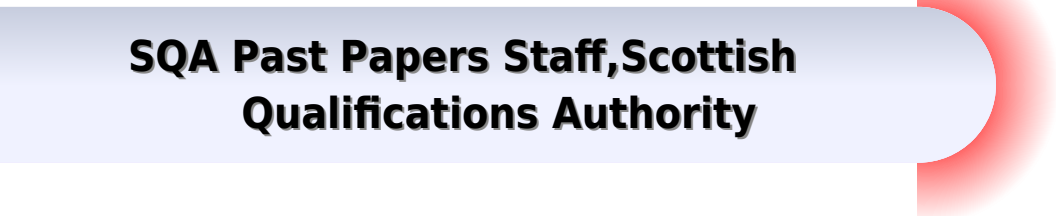


What is the relative atomic mass (A, value) of Neon?

- A) 20.18
 B) 20.28
 C) 20.10
 D) 20.22

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