

CHAPTER 1: CHEMICAL REACTIONS AND EQUATIONS

- 1. (c) 4 and 2
- 2. (d) double displacement reaction
- 3t. (a) only (i)
- 4. (b) H₂, CuO
- (a) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of Assertion (A).
- 6. (c) Assertion (A) is true but Reason (R) is false.
- 7. (i) When zinc metal is added to a solution of copper sulphate, a chemical reaction occurs where zinc displaces copper from the copper sulphate solution, forming zinc sulphate and solid copper metal. As a result, the blue colour of the copper sulphate solution fades, and a reddish-brown deposit of copper metal forms on the zinc.
 - (ii) When magnesium ribbon reacts with dilute hydrochloric acid, magnesium chloride salt will be formed and hydrogen gas is liberated. The reaction is exothermic, so the solution becomes warmer.
 - (iii) Yes, a displacement reaction is also a redox reaction because it involves the transfer of electrons where one substance is oxidised and another is reduced.
 - (iv) Galvanisation, electroplating and extraction of metals are some applications of displacement reactions in everyday life or industry.
- 8. $3\text{Fe}(s) + 4\text{H}_2\text{O}(g) \longrightarrow \text{Fe}_3\text{O}_4(s) + 4\text{H}_2(g)$ then steam steam (6.10) exists Hydrogen
- The given reaction is a redox reaction. Oxidising agent = Fe₃D₈ Reducing agent = At
- 10. (i) Decomposition reaction
 - (ii) Displacement reaction
 - (iii) Double displacement reaction
 - (iv) Combination Reaction
- 11. The salt is ferrous sulphate (FeSO₄ · 7H₂O)

Exothermic reaction is the reaction in which heat is released.

Endothermic reaction is the reaction in which heat is absorbed.

- Heating of coal in air to form carbon dioxide → Exothermic reaction.
- (ii) Heating of timestone in a time kiln to form quick time → Endothermic reaction.
- (i) The reaction will take place in presence of sunlight.

This reaction is used in black and white photography.

- (ii) (a) Displacement reaction because Cl₂ is displacing l₂ from KI solution.
 - (b) Combination reaction because K combines with Cl₂ to form a single product. KCl.

	Reaction	Substance		
		Oxidised	reduced	
(0)	$Z_{\uparrow O_{(n)}} + E_{(n)} \xrightarrow{Oxidised} Z_{\uparrow O_{(n)}} + E_{O(g)}$ reduced	С	ZnO	
(10)	Oxidised 4 Na _(S) + O _{2(S)} → 2 Na ₂ O _(S) reduced	Na	02	
(=)	CuO _(s) + H ₂ (s) x Cu(s) + H ₂ O(0)	H ₂	CuO	

- 15. (i) X-copper (Cu), Process Corrosion
 - (ii) Green coating Basic copper carbonate [CuCO₃·Cu(OH)₂]
 - (iii) Painting and oiling.
- (i) Decomposition reaction: Reaction in which a compound splits up into two or more simpler substances.
 - (ii) Oxidation reaction: The reaction in which oxygen is added to a substance or hydrogen is removed from a substance.
 - (iii) Displacement reaction: Reaction in which a more reactive element displaces a less reactive element from its compound.
 - (iv) Displacement reaction: Reaction in which a more reactive element displaces a less reactive element from its compound.
 - (v) Combination reaction: Reaction in which two or more substances combine to form a single substance.
- The process in which new substances with new properties are formed by the rearrangement of atoms is known as chemical reaction.

Four observations which helps us to determine that a chemical reaction has taken place are as follows:

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