

## ACIDS, BASES, & THE pH SCALE

Name: **Answer Key**

Date: \_\_\_\_\_

Pd: \_\_\_\_\_

### PART 1: ACIDS AND BASES

| Comparing Acids and Bases                                                                                                                                                                                                                                                                           |                                                                                                                                                                                                                                                                                                                                                                             |
|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| Acids                                                                                                                                                                                                                                                                                               | Bases (Alkalines)                                                                                                                                                                                                                                                                                                                                                           |
| <ul style="list-style-type: none"><li>• <b>Sour</b> taste, leaves <b>burning</b> sensation</li><li>• <b>Increase</b> hydrogen ions (<math>H^+</math>) in solution</li><li>• Example: <b>HCl</b>     <b><math>H^+ + Cl^-</math></b></li><li>• Turns <b>litmus</b> (pH) paper to <b>red</b></li></ul> | <ul style="list-style-type: none"><li>• Sharp <b>bitter</b> taste, feels <b>slippery</b> to touch</li><li>• <b>Decrease</b> hydrogen ions (<math>H^+</math>) in solution<br/>(<b>Increase</b> hydroxide ions (<math>OH^-</math>) in solution)</li><li>• Example: <b>NaOH</b>     <b><math>Na^+ + OH^-</math></b></li><li>• Turns litmus (pH) paper to <b>blue</b></li></ul> |

### PART 2: THE pH SCALE

- The **pH Scale** is a **measurement** of how acidic or how basic a solution is.
- The scale goes from **0** to **14**.
- **Compounds** are **acidic** if they have a pH **lower** than **7**.
- Compounds with a pH **higher** than **7** are **basic** or **alkaline**.
- Compounds with a pH **of 7**, are **neutral**.

0                      1           2           3           4           5           6           7           8           9           10           11           12           13                      14

**Increased** concentration of  **$H^+$**  ions  
**Decreased** concentration of  **$OH^-$**  ions

**Decreased** concentration of  **$H^+$**  ions  
**Increased** concentration of  **$OH^-$**  ions

pH below 7 = **acids**

pH at 7 = **neutral**  
(**pure water**)

pH above 7 = **bases**

### PART 3: BUFFERS

- A **buffer** is a **weak acid** or **base** that **prevents** sudden **changes** in pH.
  - Maintains **homeostasis**
- A buffer works by **binding** to  $H^+$  ions **releasing**  $H^+$  ions when the solution pH changes.
- Buffers **bind** to  $H^+$  ions when the **concentration** of  $H^+$  ions in solution **increases**.
- **NaOH**      **$Na^+ + OH^-$**  (The  $OH^-$  will bind to  $H^+$  ions to form  $H_2O$  – **neutralization**)
- Buffers **release**  $H^+$  ions when the **concentration** of  $H^+$  ions in solution **decreases**.  
**HCl**      **$H^+ + Cl^-$**

# 20 Acids And Bases Answer Key

**John Wilkinson**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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*General Chemistry* Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition. Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value.

**Physical Chemistry for Engineering and Applied Sciences** Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto. Designed to be as rigorous as compatible with a first year student's ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed. In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking. Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics. Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems. Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications. Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus. An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations.

*Chemistry (Teacher Guide)* Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter helps each student stay on schedule and be organized and is their source of accountability along the way. With that in mind this guide provides additional help through the laboratory exercises as well as lessons quizzes and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study.

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1. Understanding the eBook 20 Acids And Bases Answer Key
  - The Rise of Digital Reading 20 Acids And Bases Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying 20 Acids And Bases Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 20 Acids And Bases Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from 20 Acids And Bases Answer Key

- Personalized Recommendations
- 20 Acids And Bases Answer Key User Reviews and Ratings
- 20 Acids And Bases Answer Key and Bestseller Lists
- 5. Accessing 20 Acids And Bases Answer Key Free and Paid eBooks
  - 20 Acids And Bases Answer Key Public Domain eBooks
  - 20 Acids And Bases Answer Key eBook Subscription Services
  - 20 Acids And Bases Answer Key Budget-Friendly Options
- 6. Navigating 20 Acids And Bases Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - 20 Acids And Bases Answer Key Compatibility with Devices
  - 20 Acids And Bases Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 20 Acids And Bases Answer Key
  - Highlighting and Note-Taking 20 Acids And Bases Answer Key
  - Interactive Elements 20 Acids And Bases Answer Key
- 8. Staying Engaged with 20 Acids And Bases Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 20 Acids And Bases Answer Key
- 9. Balancing eBooks and Physical Books 20 Acids And Bases Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 20 Acids And Bases Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 20 Acids And Bases Answer Key
  - Setting Reading Goals 20 Acids And Bases Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 20 Acids And Bases Answer Key

- Fact-Checking eBook Content of 20 Acids And Bases Answer Key
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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