Exam-style questions and sample answers have been written by the authors. In examinations, the way marks are awarded may be different.

# > Coursebook answers

## Chapter 28

#### Science in context guidance

- Students should already be familiar with a number of wave behaviours such as reflection, refraction and diffraction. These are all phenomena that occur with light, but also with other waves such as water waves.
- Waves can also be transmitted and absorbed. It is worth asking students if they can explain wave phenomena using particles instead of waves. They may come up with some interesting ideas?

#### Self-assessment questions

- 1 energy  $E = hf = 6.63 \times 10^{-10} \times 1.0 \times 10^{20}$ = 6.63 × 10<sup>-0</sup> J = 6.6 × 10<sup>-0</sup> J
- 2 energy  $E = \frac{h_T}{4}$ For red light,  $E = \frac{6.63 \times 10^{-14} \times 3.0 \times 10^4}{700 \times 10^{-8}}$   $= 2.8 \times 10^{-19} \text{ J}$ For violet light,  $E = \frac{6.63 \times 10^{-14} \times 3.0 \times 10^4}{400 \times 10^{-8}}$  $= 5.0 \times 10^{-19} \text{ J}$
- 3 Rearrange energy E = \(\frac{k}{\chi}\) to give

wavelength  $\lambda = \frac{L}{E}$ 

- $\lambda = \frac{6.63 \times 10^{-54} \times 3.0 \times 10^{6}}{10^{-12}} \approx 2.0 \times 10^{-15} \text{ ns.}$
- b  $\lambda = \frac{6.63 \times 10^{-14} \times 3.0 \times 10^{6}}{10^{-13}} \approx 2.0 \times 10^{-16} \text{ m}.$
- $\lambda = \frac{6.63 \times 10^{-14} \times 3.0 \times 10^{4}}{\text{ultraviolet}} \approx 2.0 \times 10^{-7} \text{ m},$
- d  $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^{6}}{10^{-20}} \approx 2.0 \times 10^{-6} \text{ m},$ infrared
- $\lambda = \frac{6.63 \times 10^{-34} \times 3.0 \times 10^{8}}{10^{-29}} \approx 2.0 \text{ m, radio}$

- 4 energy of one photon,  $E = \frac{hc}{\lambda}$ =  $\frac{6.63 \times 10^{-34} \times 3.0 \times 10^{6}}{6.48 \times 10^{-2}} = 3.07 \times 10^{-13} \text{ J}$ number of photons per second =
  - power of laser =  $\frac{0.0010}{3.07 \times 10^{-18}} = 3.26 \times 10^{13} \text{ s}^{-1} \approx 3.3 \times 10^{13} \text{ s}^{-1}$
- 5 energy change, W = 1.2 eV = QV = 1.6 × 10<sup>-15</sup> × 1.2 = 1.92 × 10<sup>-15</sup> J ≈ 1.9 × 10<sup>-15</sup> J energy in eV = 1.2 eV
- 6 energy in joules, E = hf'  $1 \text{ eV} = 1.60 \times 10^{-19} \text{ J}$  $E = \frac{6.63 \times 10^{-14} \times 3.0 \times 10^{19}}{1.60 \times 10^{-19}} = 12.400 \text{ eV} \approx 12 \text{ keV}$
- 7 Rearrange energy  $E = \frac{hc}{\lambda}$  to give wavelength  $\lambda$   $= \frac{hc}{E} = \frac{6.63 \times 10^{-24} \times 3.0 \times 10^{4}}{10 \times 1.6 \times 10^{-24}} \approx 1.24 \times 10^{-2} \text{ m},$ ultraviolet
- energy gained = work done on proton by potential difference energy gained = QV = 1.60 × 10<sup>-13</sup> × 1500 = 2.4 × 10<sup>-16</sup> J
  - b kinetic energy  $E_k = \frac{1}{2} mr^2$  so speed  $r = \sqrt{\frac{2E_k}{mr}} = \sqrt{\frac{2 \times 2.4 \times 10^{-14}}{1.67 \times 10^{-23}}} = 5.3 \times 10^6 \text{ m/s}^{-1}$
- 9 Rearrange e V = \(\frac{L}{A}\) to give Planck's constant,
  - $h = \frac{cP.4}{c}$ For infrared LED,  $h_c = \frac{1.6 \times 10^{-19} \times 1.35 \times 910 \times 10^{-9}}{3.0 \times 10^9} = 6.55 \times 10^{e34} \text{ J s}$

## **Answers For Learntci Chapter 28**

**Christian G. Meyer** 

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