

## Question 1

1(a)(iii) Difference: With powdered  $\text{CaCO}_3$  the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.  
Similarity: The total volume of  $\text{CO}_2$  produced is the same in both cases since the same mass (and therefore number of moles) of  $\text{CaCO}_3$  was used in both experiments.



1(a)(v) Moles of  $\text{CaCO}_3$  used =  $(1.8/100)$  mol. = 0.018 mol.  
From the balanced equation for the reaction, 1 mole  $\text{CaCO}_3$  produces 1 mol  $\text{CO}_2$ .  
Hence volume of gas produced at R.T.P. =  $(0.018 \times 24000)$   $\text{cm}^3$  = 432  $\text{cm}^3$ .

1(a)(vi) There were impurities in the  $\text{CaCO}_3$  used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i)  $\text{Fe}^{3+}_{(aq)}$  present.

1(b)(ii) Magnesium ribbon decreases in size/dissolves.  
Solution becomes pale green.



1(b)(iv)  $\text{SO}_4^{2-}_{(aq)}$ ,  $\text{SO}_3^{2-}_{(aq)}$  and  $\text{CO}_3^{2-}_{(aq)}$  absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

# Answers For January 2014 Chemistry

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