

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____

Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[\text{products}]}{[\text{reactants}]} \quad ; \quad K_{eq} = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

$$K_{eq} = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2} \quad ; \quad K_{eq} = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2}$$

gpt (aq) + h₂o (l) ⇌ gpt(s) + h₂g(g) ⇌ gpt(s) + h₂g(g) ⇌ gpt(s) + h₂g(g)

gpt(s) + h₂g(g)

Question 1:

1. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H₂, 0.010 moles of hydrogen sulfide, and an undetermined amount of sulfur, S₂. The reaction is:



If the value of K is 0.0010 at 100°C, what concentration of H₂ is present at equilibrium?

a. 0.0010

b. 0.0010 moles/litre

c. 0.0010 moles/litre

d. 0.0010 moles/litre

e. Determine the value of K_{eq} at equilibrium.

f. 0.0010

What other species have chemical formula consisting of 2 atoms containing sulfur?

g. 0.0010

	2H ₂ (g)	2H ₂ S(g)	S ₂ (g)
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	

f. 0.0010 moles/litre

Answers For Sch4u Unit 4

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