

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat of decomposition of a substance is determined by the following thermodynamic equation:



- a. Describe how the value of ΔH_{sub} is being interpreted in the above decomposition.
- b. What is the overall ΔH value for the reaction shown in the decomposition equation?

Write the balanced equation for the reaction below. Show all given values in the space provided.

1. $\Delta H_{\text{sub}} = 20.0 \text{ kJ/mol}$ and $\Delta H_{\text{decomp}} = 10.0 \text{ kJ/mol}$ for the reaction of a substance to form a product.

2. A substance is reacting with O_2 to form a product. The reaction is exothermic, and the standard heat of formation of the product is -10.0 kJ/mol .

3. The product is a molecular compound, and the standard heat of formation of the product is -10.0 kJ/mol .

- a. What is the value of ΔH_{sub} for this reaction?

Answers For Reaction Energy Section 1

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