

8. Buffer, passing through the, is shown within the Earth's upper. Buffer is composed and the composition of these all according to the.

12 samples of Buffer from a sampling were analysed to give the following composition of samples.

| sample | composition (%) |
|--------|-----------------|
| 1%     | 0.001           |
| 1%     | 0.01            |
| 1%     | 0.02            |

- (a) Define the term relative atomic mass.

*The relative atomic mass of an element is the average mass of the atoms of the element, taking into account the relative abundance of the isotopes.*

- (b) Calculate the relative atomic mass of the sample of buffer.

Give your answer to three decimal places.

$$\left( \frac{1}{100} \times 1.001 \right) + \left( \frac{1}{100} \times 1.01 \right) + \left( \frac{1}{100} \times 1.02 \right)$$

*1.010*

answer = *1.010* (3) (3)

- (c) Some students, on using their relative atomic mass, deduced that elements were made up of tiny particles called atoms which could not be divided. However, scientists know of the existence of sub-atomic particles in atoms and in ions.

Complete the table to show the number of sub-atomic particles in the  $^{12}\text{C}$  atom and  $^{12}\text{C}^{2+}$  ion.

|                      | protons  | neutrons | electrons |
|----------------------|----------|----------|-----------|
| $^{12}\text{C}$      | <i>6</i> | <i>6</i> | <i>6</i>  |
| $^{12}\text{C}^{2+}$ | <i>6</i> | <i>6</i> | <i>4</i>  |



# Answer To F321 Chemistry May Series

**Tao Wei**



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