Analytical Chemistry MCQs: TYBSC SEM III

 The solubility product K_{sp} of sparingly soluble Ag₂CrO₄ is 4*10⁻¹². The molar solubility of the salt is

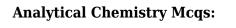
- (a) 1.0*10* moles/lit
- (b) 2.0*104 moles/lit
- (c) 1.0*10-5 moles/lit
- (d) 1.2*10-12 moles/lit

 Solubility product of AgCl is 1.44*10⁻⁴ moles/lit at 373K. The solubility of the AgCl is boiling of water will be....

- (a) 0.72*104 moles/lit
- (b) 0.72*10-2 moles/lit
- (c) 1.20*10-2 moles/lit
- (d)1.20*10⁻⁴ moles/lit
- 10. Precipitation takes place when the ionic product...
- (a) Is zero
- (b) Exceeds the solubilty product
- (c) Equal to solubilty product
- (d) Is less than the solubility product
- 11. The electomotive force is expressed In the unit of
- (a) Ohm
- (b) Volt
- (c) Ampere
- (d) Colulomb
- 12. If the cell can provide electrical energy to an external system, it is called as
- (a) Galvanic cell
- (b) Electrolytic cell
- (c) Concentration cell
- (d) Chemical cell
- The current per unit area of electrode surface is known as....
- (a) Current efficiency
- (b) Current yield
- (e) Current density
- (d) None of these

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 - The Rise of Digital Reading Analytical Chemistry Mcgs
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Analytical Chemistry Mcqs
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analytical Chemistry Mcqs
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Analytical Chemistry Mcgs
 - Personalized Recommendations
 - Analytical Chemistry Mcgs User Reviews and Ratings
 - Analytical Chemistry Mcgs and Bestseller Lists
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Utilizing eBooks for Skill Development

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