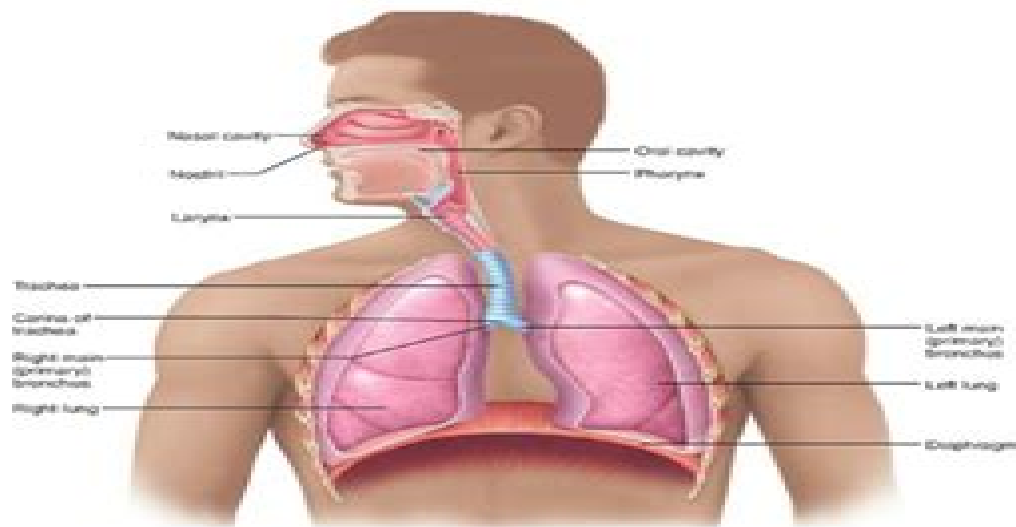


Chapter 13: The Respiratory System

I. Overview of the Respiratory System

- A. Major respiratory organs (See Figure 13.1 in text) *Be able to label the following on the figure below: nasal cavity, oral cavity, pharynx, larynx, trachea, bronchi, lungs, diaphragm.*



1. Pharynx
 - a. 3 sections are: **nasopharynx**, oropharynx and laryngopharynx
 - b. Shared by respiratory and **digestive** systems.
2. Trachea
 - a. Airway connecting the **trachea** with two bronchi that lead into the **lungs**
 - b. Fairly rigid structure because walls are reinforced with C-shaped rings of **hyaline** cartilage
 - c. Epithelium found lining the trachea is called **pseudostratified columnar epithelial tissue**
 - d. What is the function of this epithelium? **protects the lungs from irritants**

Anatomy Physiology Chapter 13 Respiratory System Answers

Valerie C. Scanlon, Tina Sanders



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monoxide co determine the mass of iodine i₂ which could be produced b if in the above situation only 0.160 moles of iodine i₂ was produced i what mass of iodine was produced

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ 4 $\text{mol Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $26.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web 7 Na_2CO_3 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of H_2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is

complete

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

reacted with 8.6 g of H_2

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reaction

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which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate
what is the

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80.0 grams of iodine pentoxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced
80 g I_2O_5 1 mol I_2O_5 1 mol I_2 $\times$ 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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give a percentage
 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$ label 3.7.3