

Tutoring questions: 17/02/14

Name:

1. Solid $\text{Al}(\text{NO}_3)_3$ is added to distilled water to produce a solution in which the concentration of nitrate, $[\text{NO}_3^-]$, is 0.10 M . What is the concentration of aluminum ion, $[\text{Al}^{3+}]$, in this solution?

- ☐ A. 0.030 M
- ☐ B. 0.033 M
- ☐ C. 0.066 M
- ☐ D. 0.10 M
- ☐ E. 0.30 M

2. A 0.20 mol sample of $\text{MgCl}_2(\text{s})$ and a 0.50 mol sample of $\text{KCl}(\text{s})$ are dissolved in water and diluted to 500 mL . What is the concentration of Cl^- in the solution?

- ☐ A. 0.15 M
- ☐ B. 0.30 M
- ☐ C. 0.50 M
- ☐ D. 0.60 M
- ☐ E. 1.0 M

A student uses visible spectrophotometry to determine the concentration of $\text{CoCl}_2(\text{aq})$ in a sample solution. First the student prepares a set of $\text{CoCl}_2(\text{aq})$ solutions of known concentrations. Then the student uses a spectrophotometer to determine the absorbance of each of the standard solutions at a wavelength of 510 nm and

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