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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$ (B) Removing a small amount of $BaO(s)$ X (C) Adding He gas to the vessel X (D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

2009 Ap Chemistry Multiple Choice

John T. Moore, Richard H. Langley



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using dimensional analysis according to first principles but the math is purposely kept at a low level and is used as a means of understanding the concepts The appendices provide a math review and a glossary of terms Carefully designed as an easy to read textbook and a practical reference this book emphasizes learning micro nanotechnology vocabulary concepts and applications from first principles and from a multi disciplinary point of view This makes it suitable for one and two semester courses as well as a reference for professionals in the field *Proceedings of the International Conference on Social Science and Education (ICoeSSE 2023)* Jhoni Warmansyah,Artha Nesa Chandra,M. Haviz,Isra Nurmai Yenti,Safrizal Safrizal,Hadiyati Idrus,2023-11-16 This is an open access book We would like to invite you to The 1st International Conference Social Science and Education ICoESSE 2023 will be conducted on September 6 7 2023 at the Universitas Islam Negeri Mahmud Yunus Batusangkar West Sumatera Indonesia The aim of the Conference is to bring together scholars students researcher and administrators from different countries and to discuss theoretical and practical issues in the fields of Science and Technology Your prospective valuable contributions on this field will be evaluated by the Scientific Committee and the ones approved to be presented will also be published in the Proceedings and selected will be publish in Journal

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