

	Bonds	C-O	C=O	C≡O	C-H	H-H	O-H
D (kJ/mol)		358	799	1072	413	436	463

A) +276 B) -276 C) +735 D) -735 E) -116

44. Which phase change results in the release of energy?

A) $\text{H}_2\text{O(g)} \rightarrow \text{H}_2\text{O(l)}$ B) $\text{H}_2\text{O(l)} \rightarrow \text{H}_2\text{O(g)}$ C) $\text{H}_2\text{O(s)} \rightarrow \text{H}_2\text{O(l)}$ D) $\text{H}_2\text{O(s)} \rightarrow \text{H}_2\text{O(g)}$

45. If the ΔH for a reaction is positive, which of the following statements is true?

A) The kinetic energy is decreasing in the system.
 B) The potential energy is decreasing in the system.
 C) The products have less potential energy than the reactants.
 D) The reactants have less potential energy than the products.

46. A temporary structure with a high potential energy that has bonds half formed and half broken is:

A) a homogeneous catalyst B) an activated complex C) radical
 D) an elementary process E) a heterogeneous catalyst

47. Consider the following reactions at 25°C.



Which of the following statements best describes the relative rates of the two reactions?

A) I and II are both fast. C) I is faster than II
 B) I and II are both slow. D) II is faster than I

48. Which of the following describes the energy of colliding particles as reacting molecules approach each other?

	KE	PE
A)	decreases	increases
B)	increases	decreases
C)	decreases	remains constant
D)	remains constant	increases

49. The rate of a chemical reaction may depend upon which factors?

I. concentration of reactants II. nature of reactants III. temperature
 A) I and II B) I and III C) II and III D) I, II, and III

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Aditya Jangid, Arshdeep Kaur



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