

Unit 9 Progress Check: FRQ

1. For parts of the free-response question that require calculations, clearly show the method used and the steps involved in arriving at your answers. You must show your work to receive credit for your answer. Examples and equations may be included in your answers where appropriate.

Answer the following questions related to H_2O .

Substance	ΔG°_f at 298 K (kJ/mol)
$\text{H}_2\text{O}(l)$	-237.2
$\text{H}_2\text{O}(g)$	-228.4

(a) Using the information in the table above, determine the value of ΔG° at 298 K for the process represented by the equation $\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{O}(g)$.



Please respond on separate paper, following directions from your teacher.

(b) Considering your answer to part (a), indicate whether the process is thermodynamically favorable at 298 K. Justify your answer.



Please respond on separate paper, following directions from your teacher.

(c) Considering your answer to part (b), explain why $\text{H}_2\text{O}(l)$ has a measurable equilibrium vapor pressure at 298 K.



Please respond on separate paper, following directions from your teacher.

Water vapor can be produced in two different processes, as represented below.

Process 1	Process 2
$\text{H}_2\text{O}(s) \rightleftharpoons \text{H}_2\text{O}(g)$	$\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{O}(g)$

(d) In terms of concepts of entropy and the particle-level structure of the different phases of water, explain why the change in entropy, ΔS , is greater for process 1 than for process 2.

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