

CHAPTER 14 REVIEW*Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Name the following compounds as acids:

_____ sulfuric acid _____	a. H_2SO_4
_____ sulfurous acid _____	b. H_2SO_3
_____ hydrosulfuric acid _____	c. H_2S
_____ perchloric acid _____	d. HClO_4
_____ hydrocyanic acid _____	e. hydrogen cyanide

- 2.**
- _____
- H_2S
- _____ Which (if any) of the acids mentioned in item
- 1**
- are binary acids?

- 3.**
- Write formulas for the following acids:

_____ HNO_2 _____	a. nitrous acid
_____ HBr _____	b. hydrobromic acid
_____ H_3PO_4 _____	c. phosphoric acid
_____ CH_3COOH _____	d. acetic acid
_____ HClO _____	e. hypochlorous acid

- 4.**
- Calcium selenate has the formula
- CaSeO_4
- .

_____ H_2SeO_4 _____	a. What is the formula for selenic acid?
_____ H_2SeO_3 _____	b. What is the formula for selenous acid?

- 5.**
- Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

- 6.**
- Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



Describing Acids And Bases Section Review

Karen Wilson, Adam F. Lee



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