

# DESCRIBING CHEMICAL REACTIONS

## Chemical Reactions

- Chemical reactions are fundamental processes that occur in everything
- A chemical reaction involves the transformation of one or more substances into new substances with different properties

## Coefficients

- A coefficient of a chemical formula tells you how many molecules are involved in the reaction



- The "2" in front of the H<sub>2</sub>O tells you that you have two molecules of water
- How many H are there? 4
- How many O are there? 2

## Law of Conservation of Mass

- States that matter cannot be created or destroyed
- The total mass of the reactants must equal the total mass of the products
- Because atoms are the smallest unit of an element, you must have the same number of each on the reactant side and the product side

## Chemical Formulas

- To represent chemical reactions, we use chemical equations
- Chemical equations are symbolic representations of the reactants and products in the reaction
- Chemical formulas describe the elements and number of atoms of each element involved

## Chemical Formulas

- Chemical symbols contain 1 or 2 letters and the first letter is always capitalized
- Subscripts are the small numbers to the bottom right of the chemical symbol

## Steps for Balancing Equations

1. Draw a T-chart on both sides
2. List all elements on the left side of the chart
3. List the number of atoms of each element on both sides
4. Adjust the coefficients until you have the same number of atoms of each element on both sides

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**Suzanne Slade**



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Prentice Hall Science Explorer: Chemical Interactions Pearson/Prentice Hall, Michael J. Padilla, 2004-10      **Focus on Physical Science California Edition** Michael J. Padilla, 2007      Current Index to Journals in Education , 1989

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