

Time allowed: 2:40 Hours

Total Marks Sections B and C:

NOTE: Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separate provided answer book. Write your answers neatly and legibly. 09280208

SECTION - B (Marks 33)

(11 × 3 =

Q. 2 Attempt any ELEVEN parts. All parts carry equal marks.

- (i) Briefly explain the following. Also give one example of each
 - (a) Ion
 - (b) Environmental Chemistry
 - (c) Atomic Mass
- (ii) Compare compounds and mixtures. Also give examples.
- (iii) There are three isotopes of hydrogen. Draw their diagrams.
- (iv) Distinguish between 'shell' and 'subshell'.
- (v) State periodic law. Also write the names of groups in s block of the periodic table.
- (vi) Show the valence shell electronic configuration for the following:
 - a. Alkali metals s^1
 - b. Nobel gases s^2, p^6
 - c. Halogens s^2, p^5
- (vii) Determine the location of Cr^{4+} in periodic table. $6 = 2, 2 =$
- (viii) Briefly explain the following terms: (Any two)
 - a. Melting
 - b. Freezing
 - c. Boiling
- (ix) Find the number of electrons in valence shell in the following:
 - a. Neon
 - b. Helium
 - c. Hydrogen
- (x) Calculate the molarity of a solution that contains 40g urea (NH_2CONH_2) dissolved in 500cm³ of solution
 $14 + 2 \times 1 + 12 + 16 + 14 + 16 + 1 \times 2$
- (xi) The atomic number of an element is 20 and mass number is 40. How many Protons, Neutrons and Electrons does an atom of this element have?
- (xii) Enlist the possible uses of an electrolytic cell.
- (xiii) Find oxidation state of Nitrogen in any two of given compounds:
 - a. N_2O_5
 - b. HNO_3
 - c. NO_2
- (xiv) Magnesium reacts with steam, but calcium reacts with cold water. Show this diversity in reactivity by chemical equations.
- (xv) Why fluorine has the highest oxidizing power? Explain briefly.

SECTION - C (Marks 20)

(2 × 10 = 20)

Note: Attempt any TWO questions. All questions carry equal marks.

- Q. 3 a. What is meant by evaporation? The water level in an aquarium decreases slowly, even though the tank doesn't leak. Interpret
- b. Compare the effects of Al_2O_3 and Fe_2O_3 formation on their parent metals. (Al and Fe)

- Q. 4 a. KCl and MgF_2 are two ionic compounds. Discuss the bond formation in each of these. Also draw electron dot and cross models.
- b. Briefly discuss the following types of solutions. also give one example of each:

	Solute	Solvent
I.	Solids	Liquids
II.	Solids	Solids

- Q. 5 a. What are unsaturated, saturated and supersaturated solutions? Also give their comparison.
- b. Describe the formation of anions in Bromine and Phosphorous ions.

SUPPLEMENTARY TABLE

— 15A-1 2308 HA — 09280208

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40



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 - The Rise of Digital Reading 1class Board Chemistry Paper 2hbse
 - Advantages of eBooks Over Traditional Books
2. Identifying 1class Board Chemistry Paper 2hbse
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1class Board Chemistry Paper 2hbse
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1class Board Chemistry Paper 2hbse

- Personalized Recommendations
- 1class Board Chemistry Paper 2hbse User Reviews and Ratings
- 1class Board Chemistry Paper 2hbse and Bestseller Lists
- 5. Accessing 1class Board Chemistry Paper 2hbse Free and Paid eBooks
 - 1class Board Chemistry Paper 2hbse Public Domain eBooks
 - 1class Board Chemistry Paper 2hbse eBook Subscription Services
 - 1class Board Chemistry Paper 2hbse Budget-Friendly Options
- 6. Navigating 1class Board Chemistry Paper 2hbse eBook Formats
 - ePub, PDF, MOBI, and More
 - 1class Board Chemistry Paper 2hbse Compatibility with Devices
 - 1class Board Chemistry Paper 2hbse Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1class Board Chemistry Paper 2hbse
 - Highlighting and Note-Taking 1class Board Chemistry Paper 2hbse
 - Interactive Elements 1class Board Chemistry Paper 2hbse
- 8. Staying Engaged with 1class Board Chemistry Paper 2hbse
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1class Board Chemistry Paper 2hbse
- 9. Balancing eBooks and Physical Books 1class Board Chemistry Paper 2hbse
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1class Board Chemistry Paper 2hbse
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1class Board Chemistry Paper 2hbse
 - Setting Reading Goals 1class Board Chemistry Paper 2hbse
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1class Board Chemistry Paper 2hbse

- Fact-Checking eBook Content of 1class Board Chemistry Paper 2hbse
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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