

POGIL 4 Key

Coulombic Attraction

What variables will affect the force of attraction between charged particles?

Coulombic attraction is the attraction between oppositely charged particles. For example, the protons in the nucleus of an atom have attraction for the electrons surrounding the nucleus. This is because the protons are positive and the electrons are negative. The attractive force can be weak or strong.

Explore the strength of attraction between protons and electrons in various atomic structures below.

Distance and Force of attraction

		Force of Attraction (Newtons)
A		2.30×10^{-8}
B		0.58×10^{-8}
C		0.26×10^{-8}

- What subatomic particles do these symbols represent in the drawing above:



Proton



Electron

- What are the independent and dependent variables in the data?

*Charge and distance are independent
Force of attraction is dependent*

- What relationship do you observe between the independent and dependent variables in this case?

Force of attraction between two particles is directly proportional to charges of particles and inversely proportional to the distance.

- If the distance between a proton and electron were 0.50 nm, would you expect the force of attraction to be greater than or less than $0.26 \times 10^{-8}\text{N}$? Explain.

The force of attraction would be weaker than $0.26 \times 10^{-8}\text{N}$. It is inversely proportional to the distance. Greater distance = weaker force.

- If **two** protons were 0.10 nm away from one electron, would you expect the force of attraction to be greater than or less than $2.30 \times 10^{-8}\text{N}$? Explain.

I would expect the force of attraction to be greater, since the charge of two protons is bigger than the charge of one.

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Katrin Zwirgmaier



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