

change in pH by adding NaOH would be 1.85.

(C) In solution, $[\text{Na}^+] = 0.1 \text{ M}$, $[\text{NH}_3] = 0.1 \text{ M}$, $[\text{OH}^-] = 0.2 \text{ M}$.

(D) on addition of OH^- , K_b of NH_3 does not change.

Calculate the pH of solution obtained by mixing 10 mL of 0.1 M HCl and 40 mL of 0.2 M H_2SO_4 . Take $\log 3.4 = 0.53$

The pH of 0.2 M $\text{B}(\text{OH})_2$ solution is 1.13×10^4 , determine x
($K_{b1} = 2 \times 10^{-5}$; $K_{b2} = 4 \times 10^{-11}$, $\log 2 = 0.3$)

The $[\text{S}^{2-}]$ in a solution in which 0.1 M HCl & 0.1 M H_2S is present, given that $K_{a1}(\text{H}_2\text{S}) = 10^{-7}$,
 $K_{a2}(\text{HS}^-) = 10^{-14}$ is 10^{-x} , determine x

What is the pH of 1 M Na_2PO_4 solution at 25°C ?

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J. Gregor



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