Exercises

reaction (= N2 + 3H2 -> 2NH3

- 1. Provide an expression for Kp for reaction 1. Why does a term for C(graphite) not appear in the Kr expression for the reaction of Table 17.
- 2. Calculate the equilibrium constant for the reaction using tabulated data.

DrH = -484 by Ind

2 H2(g) + O2(g) = 2 H2O(g) at 298 K, 1 bar

Dr50 -- 89 Ks/mel

Write the equilibrium constant for this reaction in terms of the partial pressures.

3. Calculate the equilibrium constant for

 $\Delta_{CG}(Co_1) = -604.3k_{KA}$ $CaCO_3(calcite) \Longrightarrow CaO(s) + CO_2(g)$ at 298 K, 1 bar from tabulated data.

(CoCo₂): -1128.77 leg/col

9. The reaction for the formation of nitrosyl chloride

 $2NO(g) + Cl_2(g) \Longrightarrow 2NOCl(g)$

was studied at 25°C. The pressures at equilibrium were found to be

 $P_{\text{NOCI}} = 1.2 \text{ atm}$

 $P_{NO} = 5.0 \times 10^{-2} \text{ atm}$

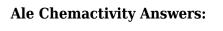
 $P_{Cl_2} = 3.0 \times 10^{-1} \text{ atm}$

- Calculate the value of Kp for this reaction at 25°C.
- If 1.0 atm of NOCI(g) was placed into an evacuated container would a 60.
- If 1.0 atm each of NOCl(g), Cl2(g) and NO(g) were placed into an evacuated ch: container would a reaction occur? If so in which direction would the reaction proceed?
- If 1.0 atm of NOCl(g), 5.0×10^{-3} atm of NO and 3.0×10^{-1} atm of Cl₂ were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

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