

## Exercises



1. Provide an expression for  $K_P$  for reaction 1. Why does a term for C(graphite) not appear in the  $K_P$  expression for the reaction of Table 1?
2. Calculate the equilibrium constant for the reaction using tabulated data.

$$\Delta_r H^\circ = -484 \text{ kJ/mol}$$

$$\Delta_r S^\circ = -89 \text{ kJ/mol}$$



Write the equilibrium constant for this reaction in terms of the partial pressures.

3. Calculate the equilibrium constant for

$$\Delta_r G_f^\circ(\text{CaO}) = -604.3 \text{ kJ/mol}$$

$$\Delta_r G_f^\circ(\text{CO}_2) = -394.36 \text{ kJ/mol}$$

from tabulated data.

$$\Delta_r G_f^\circ(\text{CaCO}_3) = -1128.77 \text{ kJ/mol}$$



9. The reaction for the formation of nitrosyl chloride



was studied at 25°C. The pressures at equilibrium were found to be

$$P_{\text{NOCl}} = 1.2 \text{ atm}$$

$$P_{\text{NO}} = 5.0 \times 10^{-2} \text{ atm}$$

$$P_{\text{Cl}_2} = 3.0 \times 10^{-1} \text{ atm}$$

- a) Calculate the value of  $K_P$  for this reaction at 25°C.
- b) If 1.0 atm of  $\text{NOCl}(\text{g})$  was placed into an evacuated container would a reaction occur?
- c) If 1.0 atm each of  $\text{NOCl}(\text{g})$ ,  $\text{Cl}_2(\text{g})$  and  $\text{NO}(\text{g})$  were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?
- d) If 1.0 atm of  $\text{NOCl}(\text{g})$ ,  $5.0 \times 10^{-3}$  atm of  $\text{NO}$  and  $3.0 \times 10^{-1}$  atm of  $\text{Cl}_2$  were placed into an evacuated container would a reaction occur? If so in which direction would the reaction proceed?

Show all calculations and explain each answer.

# Ale Chemactivity Answers

**LM Reder**



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## **Table of Contents Ale Chemactivity Answers**

1. Understanding the eBook Ale Chemactivity Answers
  - The Rise of Digital Reading Ale Chemactivity Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ale Chemactivity Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ale Chemactivity Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ale Chemactivity Answers
  - Personalized Recommendations
  - Ale Chemactivity Answers User Reviews and Ratings
  - Ale Chemactivity Answers and Bestseller Lists
5. Accessing Ale Chemactivity Answers Free and Paid eBooks

- Ale Chemactivity Answers Public Domain eBooks
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- Ale Chemactivity Answers Budget-Friendly Options
- 6. Navigating Ale Chemactivity Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ale Chemactivity Answers Compatibility with Devices
  - Ale Chemactivity Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ale Chemactivity Answers
  - Highlighting and Note-Taking Ale Chemactivity Answers
  - Interactive Elements Ale Chemactivity Answers
- 8. Staying Engaged with Ale Chemactivity Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ale Chemactivity Answers
- 9. Balancing eBooks and Physical Books Ale Chemactivity Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ale Chemactivity Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ale Chemactivity Answers
  - Setting Reading Goals Ale Chemactivity Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ale Chemactivity Answers
  - Fact-Checking eBook Content of Ale Chemactivity Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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