

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in two parts. In the first part, a known mass of a sample containing chloride ions is dissolved in water. A known volume of a silver nitrate solution is added to the sample, and the mixture is allowed to react. The precipitate is then filtered, dried, and weighed. In the second part, a known mass of a sample containing chloride ions is dissolved in water. A known volume of a silver nitrate solution is added to the sample, and the mixture is allowed to react. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

The reaction between silver ions and chloride ions is as follows:

$$\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$$

The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Procedure

The procedure for this experiment is as follows:

1. Weigh a known mass of a sample containing chloride ions.
2. Dissolve the sample in water.
3. Add a known volume of a silver nitrate solution to the sample.
4. Allow the mixture to react.
5. Filter the precipitate.
6. Dry the precipitate.
7. Weigh the precipitate.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated as follows:

$$\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of chloride ion in the sample is calculated as follows:

Analysis Of Chloride Lab Report

Michael Seilmaier



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web use your table to note down examples of persuasive writing as they appear in the article some examples have been given to start you off look back at the learn section if you need a recap