

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and sodium nitrate. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a hood to avoid exposure to the fumes of silver chloride, which can be irritating to the eyes and respiratory system. The experiment is performed in a hood to avoid exposure to the fumes of silver nitrate, which can be irritating to the eyes and respiratory system. The experiment is performed in a hood to avoid exposure to the fumes of silver chloride, which can be irritating to the eyes and respiratory system.

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Procedure

The procedure involves the following steps: 1. Weigh a sample of the sample solution. 2. Dissolve the sample in water. 3. Add silver nitrate to the solution to form a precipitate of silver chloride. 4. Filter the precipitate. 5. Dry the precipitate. 6. Weigh the dried precipitate. 7. Calculate the mass of chloride in the original sample.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

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Table of Contents Analysis Of Chloride Experiment 7 Lab Report

1. Understanding the eBook Analysis Of Chloride Experiment 7 Lab Report
 - The Rise of Digital Reading Analysis Of Chloride Experiment 7 Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Chloride Experiment 7 Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Chloride Experiment 7 Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Chloride Experiment 7 Lab Report
 - Personalized Recommendations
 - Analysis Of Chloride Experiment 7 Lab Report User Reviews and Ratings
 - Analysis Of Chloride Experiment 7 Lab Report and Bestseller Lists

5. Accessing Analysis Of Chloride Experiment 7 Lab Report Free and Paid eBooks
 - Analysis Of Chloride Experiment 7 Lab Report Public Domain eBooks
 - Analysis Of Chloride Experiment 7 Lab Report eBook Subscription Services
 - Analysis Of Chloride Experiment 7 Lab Report Budget-Friendly Options
6. Navigating Analysis Of Chloride Experiment 7 Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of Chloride Experiment 7 Lab Report Compatibility with Devices
 - Analysis Of Chloride Experiment 7 Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Chloride Experiment 7 Lab Report
 - Highlighting and Note-Taking Analysis Of Chloride Experiment 7 Lab Report
 - Interactive Elements Analysis Of Chloride Experiment 7 Lab Report
8. Staying Engaged with Analysis Of Chloride Experiment 7 Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Chloride Experiment 7 Lab Report
9. Balancing eBooks and Physical Books Analysis Of Chloride Experiment 7 Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Chloride Experiment 7 Lab Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Analysis Of Chloride Experiment 7 Lab Report
 - Setting Reading Goals Analysis Of Chloride Experiment 7 Lab Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Analysis Of Chloride Experiment 7 Lab Report
 - Fact-Checking eBook Content of Analysis Of Chloride Experiment 7 Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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