Calculations in Analytical Chemistry

The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, NA=6.022X10^23.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

| Element | Molar mass per element | Number of atoms | Mass |
|------------|------------------------------|--------------------|--------------|
| С | 12 g/mol | 6 | 72 g/mol |
| н | 1 g/mol | 12 | 12g/m |
| 0 | 16 g/mol | 6 | 96 g/mol |
| Molar Mass | | | 180 g/mol |

Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

 Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 \frac{g_{HBz}}{g_{HBz}} \times \frac{1mol_{HBz}}{122.1 \frac{g_{HBz}}{g_{HBz}}} = 0.0164 mol_{HBz}$$

To obtain the number of milimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 \frac{g_{HBz}}{g_{HBz}} \times \frac{1mmol_{HBz}}{0.1221 \frac{g_{HBz}}{g_{HBz}}} = 16.4mmol_{HBz}$$

 What is the mass in grams of Na* (22.99 g/mol) in 25.0 g of Na₂SO₄ (142.0 g/mol)?

$$mass_{Na^{+}} = 25.0 \frac{g_{Na_{1}50_{4}}}{142.0 \frac{g_{Na_{1}50_{4}}}{142.0 \frac{g_{Na_{1}50_{4}}}{1400 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{2 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{22.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{22.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{20.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{20.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}$$

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