

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.

- This increases the $[H_2]$, which will increase the rate, but has no effect on k .
- Due to Arrhenius equation, changing temperature changes the value of k .
- Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/(L · s)
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to:
 P_4 0
 H_2 1
- Write the rate law for this reaction. rate = $k[H_2]$
- Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 \text{ s}^{-1}$

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/(L · s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to:
 SO_2 2
 O_3 0
- Write the rate law for this reaction. rate = $k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 \text{ mol/L}^{-2} \text{ s}^{-1}$

8. Consider the following mechanism.

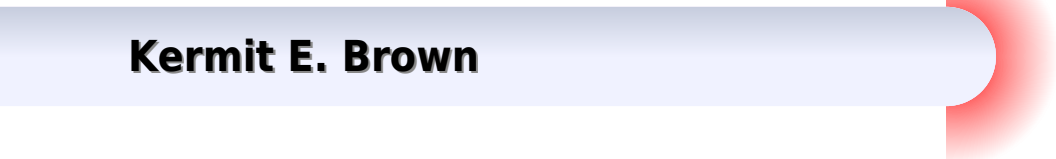
$$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$

$$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. R
- What is the order with respect to each reactant? A_2 1st; B_2 1st

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Kermit E. Brown



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Chemical kinetics Elena Burlakova, Alexander Evgen'evich Shilov, Sergey Dmitrievich Varfolomeev, Gennady Zaikov, 2005-10-28 The volume is devoted to the problem of chemical kinetics on modern level The book includes information on chemical physics of nanocomposites degradation stabilization and flammability of polymeric materials as well as free radical mechanism of oxidation of organic compounds thermostability mechanism of action of catalytical systems and inhibitors in free radical reactions in liquid and solid phase pure and applied chemistry of antioxidants synthesis and application ionic reactions effect of chemoluminescence in the processes of oxidation biodegradation and application of polymers in medicine problems of adhesion of microorganisms on the surface of materials thermo photo and hydrolytic reactions creation of new ecologically friendly flame retardants for polymers polymer composites and polymer blends as well as filled polymers

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