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AP CHEMISTRY REVIEW WORKSHEET

(Units 11&12 – Thermodynamics and Electrochemistry)

1. Which of the following must be true for a reaction that proceeds spontaneously from initial standard state conditions?

- a. $\Delta G^\circ > 0$ and $K_{eq} > 1$
- b. $\Delta G^\circ > 0$ and $K_{eq} < 1$
- c. $\Delta G^\circ < 0$ and $K_{eq} > 1$
- d. $\Delta G^\circ < 0$ and $K_{eq} < 1$
- e. $\Delta G^\circ = 0$ and $K_{eq} = 1$

2.
$$\text{N}_{2(g)} + 3 \text{H}_{2(g)} \leftrightarrow 2 \text{NH}_{3(g)}$$

The reaction indicated above is thermodynamically spontaneous at 298 K, but becomes non-spontaneous at higher temperatures. Which of the following is true at 298 K?

- a. ΔG , ΔH , and ΔS are all positive.
- b. ΔG , ΔH , and ΔS are all negative.
- c. ΔG and ΔH are negative, but ΔS is positive.
- d. ΔG , and ΔS are negative, but ΔH is positive.
- e. ΔG and ΔH are positive, but ΔS is negative.

3. A cube of ice is added to some hot water in a rigid, insulated container, which is then sealed. There is no heat exchange with the surroundings. What has happened to the total energy and the total entropy when the system reaches equilibrium?

- | | Energy | Entropy |
|----|------------------|------------------|
| a. | Remains constant | Remains constant |
| b. | Remains constant | Decreases |
| c. | Remains constant | Increases |
| d. | Decreases | Increases |
| e. | Increases | Decreases |

4. Which of the following reactions has the largest positive value of ΔS per mole of Cl_2 ?

- a. $\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightarrow 2 \text{HCl}_{(g)}$
- b. $\text{Cl}_{2(g)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{Cl}_2\text{O}_{(g)}$
- c. $\text{Mg}_{(s)} + \text{Cl}_{2(g)} \rightarrow \text{MgCl}_{2(s)}$
- d. $2 \text{NH}_4\text{Cl}_{(s)} \rightarrow 4 \text{H}_{2(g)} + \text{Cl}_{2(g)}$
- e. $\text{Cl}_{2(g)} \rightarrow 2 \text{Cl}_{(g)}$

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