

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until the chloride ions are completely precipitated as silver chloride. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on an analytical balance. 6. Calculation: The mass of chloride ions in the original solution is calculated from the mass of the precipitate.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate.

Procedure

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until the chloride ions are completely precipitated as silver chloride. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on an analytical balance. 6. Calculation: The mass of chloride ions in the original solution is calculated from the mass of the precipitate.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated from the mass of the precipitate as follows:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of chloride ion in the sample is calculated from the mass of the precipitate as follows:

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