

**Review for the Full Year General Chemistry
American Chemical Society Exam**

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$]

- (A) 0.11 J/g K (C) 17 J/g K
(B) 0.22 J/g K (D) 34 J/g K

85. What is ΔH in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of $\text{SO}_3(\text{g})$.

- (A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

- (A) 88 kJ (B) 109 kJ
(C) 175 kJ (D) 284 kJ

88. Calculate the value of ΔH (in kJ mol⁻¹) for the reaction.



Bond Energies (kJ mol⁻¹)

H-H	435
N=N	945 (in N ₂)
N-H	389

- (A) 2340 kJ of heat absorbed
(B) 213 kJ of heat absorbed
(C) 2340 kJ of heat evolved
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

(A) dissolving sugar in water

(B) expanding a gas

(C) evaporating a liquid

(D) freezing water

92. When $\text{Al}_2\text{O}_3(\text{s})$ is formed from the elements at standard conditions, the values of ΔH° and ΔG° at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

(A) ΔH , ΔS , and ΔG are positive at all temperatures.

(B) ΔH and ΔS are positive.

(C) ΔG is negative at low temperatures, positive at high temperatures.

(D) $\Delta H = \Delta S$

94. A particular chemical reaction has a negative ΔH and negative ΔS . Which statement is correct?

(A) The reaction is spontaneous at all temperatures.

(B) The reaction is nonspontaneous at all temperatures.

(C) The reaction becomes spontaneous as temperature increases.

(D) The reaction becomes spontaneous as temperature decreases.

American Chemical Society Exam Study Guide 2015

Marco Cascella



American Chemical Society Exam Study Guide 2015:

CAE Study Guide 2015 American Society of Association Executives (ASAE), 2015-02-09 The ASAE CAE preparation guide featuring 100 exam questions This is the ASAE test prep resource for the CAE exam which leads to certification as an association executive The ASAE CAE Study Guide 2015 Preparation Reference for the Certified Association Executive Exam Practice Test Bank is the only book with a practice exam based on actual questions released from the official CAE item bank This edition reflects newly updated and revised job analysis study findings from the American Society of Association Executives ASAE In this book you will find everything you need to feel confident on test day including high level summaries of all the content covered in the CAE content outline's nine domains In addition to easy to digest content sections CAE Study Guide 2015 is full of indispensable resources for anyone preparing for CAE certification The online practice exam includes 100 questions released from the official CAE exam database and testing tips will orient you toward the most successful strategies Comprehensive content summaries of all nine domains tested on the CAE exam Online practice exam consisting of 100 questions drawn from the CAE test item database plus answer key Preparation resources including recommended reading key definitions self assessment CAE study questions and more Exam day readiness tips and suggestions for creating a study plan focusing on difficult areas and reducing test anxiety Beginning with the May 2015 CAE exam the CAE Commission will implement its revised CAE test based on the updated content structure CAE Study Guide 2015 is a terrifically useful resource for anyone seeking CAE certification

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Chemistry Education and Contributions from History and Philosophy of Science Mansoor Niaz, 2015-12-23 This book explores the relationship between the content of chemistry education and the history and philosophy of science HPS framework that underlies such education It discusses the need to present an image that reflects how chemistry developed and progresses It proposes that chemistry should be taught the way it is practiced by chemists as a human enterprise at the

interface of scientific practice and HPS Finally it sets out to convince teachers to go beyond the traditional classroom practice and explore new teaching strategies The importance of HPS has been recognized for the science curriculum since the middle of the 20th century The need for teaching chemistry within a historical context is not difficult to understand as HPS is not far below the surface in any science classroom A review of the literature shows that the traditional chemistry classroom curricula and textbooks while dealing with concepts such as law theory model explanation hypothesis observation evidence and idealization generally ignore elements of the history and philosophy of science This book proposes that the conceptual understanding of chemistry requires knowledge and understanding of the history and philosophy of science Professor Niaz s book is most welcome coming at a time when there is an urgently felt need to upgrade the teaching of science The book is a huge aid for adding to the usual way presenting science as a series of mere facts also the necessary mandate to show how science is done and how science through its history and philosophy is part of the cultural development of humanity Gerald Holton Mallinckrodt Professor of Physics these are valence bond VB and molecular orbital MO theories He emphasizes the importance of having rival theories and interpretations in science and its advancement He further argues that this VB MO rivalry is still alive and together the two conceptual frames serve as the tool kit for thinking and doing chemistry in creative manners The author surveys chemistry textbooks in the light of the how the books preserve or not the balance between the two theories in describing various chemical phenomena This Talmudic approach of conceptual tension is a universal characteristic of any branch of evolving wisdom As such Mansoor s book would be of great utility for chemistry teachers to examine how can they become more effective teachers by recognizing the importance of conceptual tension Sason Shaik Saeree K and Louis P Fiedler Chair in Chemistry Director The Lise Meitner Minerva Center for Computational Quantum Chemistry The Hebrew University of Jerusalem ISRAEL **The Cambridge Handbook of Computing**

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Journal of Economic Literature ,1984 *Alcohol and Drug Counselor (ADC)*

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