

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on an analytical balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate.

Procedure:

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on an analytical balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated as follows: $\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$

For Sample 1: $\text{Mass of Cl}^- = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

The mass of chloride ion in the sample is calculated as follows: $\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$

Analysis Of Chloride Lab Report

Martha M. Makholm



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Table of Contents Analysis Of Chloride Lab Report

1. Understanding the eBook Analysis Of Chloride Lab Report
 - The Rise of Digital Reading Analysis Of Chloride Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Chloride Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Chloride Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Chloride Lab Report
 - Personalized Recommendations

- Analysis Of Chloride Lab Report User Reviews and Ratings
- Analysis Of Chloride Lab Report and Bestseller Lists
- 5. Accessing Analysis Of Chloride Lab Report Free and Paid eBooks
 - Analysis Of Chloride Lab Report Public Domain eBooks
 - Analysis Of Chloride Lab Report eBook Subscription Services
 - Analysis Of Chloride Lab Report Budget-Friendly Options
- 6. Navigating Analysis Of Chloride Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of Chloride Lab Report Compatibility with Devices
 - Analysis Of Chloride Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Chloride Lab Report
 - Highlighting and Note-Taking Analysis Of Chloride Lab Report
 - Interactive Elements Analysis Of Chloride Lab Report
- 8. Staying Engaged with Analysis Of Chloride Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Chloride Lab Report
- 9. Balancing eBooks and Physical Books Analysis Of Chloride Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Chloride Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of Chloride Lab Report
 - Setting Reading Goals Analysis Of Chloride Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Chloride Lab Report
 - Fact-Checking eBook Content of Analysis Of Chloride Lab Report

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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