

Atomic Structure Worksheet

Label the parts of an atom on the diagram below.



4. What type of charge does a proton have?

Positive (+1)

5. What type of charge does a neutron have?

Neutral (0)

6. What type of charge does an electron have?

Negative (-1)

7. Which two subatomic particles are located in the nucleus of an atom?

Protons and neutrons

8. If an atom has 35 protons in the nucleus, how many electrons will it have orbiting the nucleus?

35

9. What is the atomic number of the atom in the diagram above?

5

10. What is the atomic mass/mass number of the atom in the diagram above?

(11+11=22) *protons and neutrons added together) 22 is the atomic mass/mass number

11. How many protons are in the nucleus of an atom with an atomic number of 15?

15 *number of protons is the same as the atomic number

12. How many electrons are in the nucleus of an atom with an atomic number of 20?

20

13. How many neutrons are in the nucleus of an atom with an atomic number of 25?
(use Periodic Table for mass)

30 Neutrons. atomic mass: 54.938 ~ 55 atomic number: 25.....55-25=30

14. What is the mass number of an atom with 3 protons, 4 neutrons, and 3 electrons?

Mass number= 6.94. (Li-Lithium has an atomic number of 3 which means it has 3 protons.

15. How many neutrons are in the nucleus of an atom that has an atomic mass of 36 and an atomic number of 25?

Neutrons= 11 (atomic mass: 36 - 25 :atomic number)

Analyzing Atomic Structure Worksheet Answers

Kristy Turner



Analyzing Atomic Structure Worksheet Answers:

Problems and Problem Solving in Chemistry Education Georgios Tsaparis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **Popular Electronics** ,1982 *Science Interactions* ,1996

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