

Stoichiometry Limiting Reactant

Answers

- (1) How many grams of calcium phosphate can be produced from the reaction between 2.50 L of 0.250 M calcium chloride with an excess of phosphoric acid?



$$2.50 \text{ L CaCl}_2 \times \frac{0.250 \text{ mol CaCl}_2}{1 \text{ L CaCl}_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \times \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 64.6 \text{ g Ca}_3(\text{PO}_4)_2$$

- (2) How many milliliters of 1.50 M nitric acid is required to react with 100.0 g of cuprous oxide?



$$100 \text{ g Cu}_2\text{O} \times \frac{1 \text{ mol Cu}_2\text{O}}{143 \text{ g Cu}_2\text{O}} \times \frac{14 \text{ mol HNO}_3}{3 \text{ mol Cu}_2\text{O}} \times \frac{1000 \text{ ml HNO}_3}{1.5 \text{ mol HNO}_3} = 2180 \text{ ml HNO}_3$$

- (3) Consider the following unbalanced reaction:



Balance the reaction. 3.25 g of NH_3 is allowed to react with 3.5 g of O_2 .

- (a) Which reactant is the limiting reagent?

$$3.25 \text{ g NH}_3 \times \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 5.73 \text{ g NO}$$

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 2.63 \text{ g NO}$$

O_2 is the limiting reagent.

- (b) How many grams of NO are formed?

2.63 g of NO are formed.

- (c) How much excess reagent remains after the reaction?

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NH}_3}{5 \text{ mol O}_2} \times \frac{17.03 \text{ g NH}_3}{1 \text{ mol NH}_3} = 1.49 \text{ g NH}_3 \text{ used}$$

$$3.25 \text{ g NH}_3 - 1.49 \text{ g NH}_3 = 1.76 \text{ g NH}_3 \text{ remains.}$$

Answer Sheet To The Stoichiometry Limiting Reagent

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Barreto, Carlos Daniel Luzi, 2024-03-12 This book delves into the realm of Chemical Reaction Engineering CRE by showcasing the practical application of multiple reaction stoichiometry The authors critically assess various approaches commonly taught in undergraduate CRE courses to establish the relationships between changes in chemical species In doing so they propose an innovative conceptual alternative that is specifically tailored for undergraduate lectures The book carefully selects composition measures that effectively harness the power of stoichiometric relationships in elementary reacting systems and models which are typically covered in these courses Going beyond the basics it also offers a profound discussion on the value of chemical stoichiometry for tackling more intricate reaction systems and detailed models Moreover the book presents a simplified procedure that minimizes the reliance on complex linear algebra techniques making the book accessible to a wider range of readers

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