

OSMOSIS WORKSHEET

Name: _____

Part One: Color the correct box in the chart below:

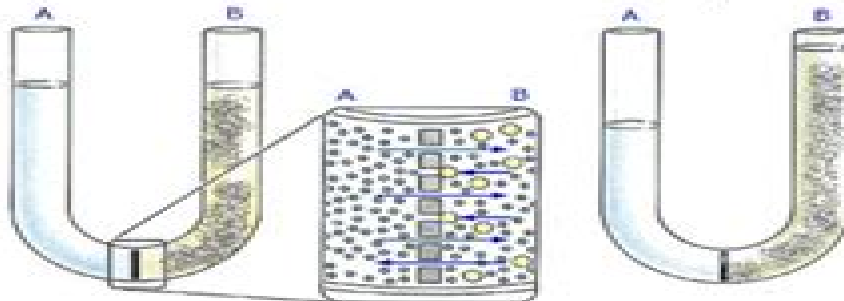
STATEMENT	Isotonic Cell(s)	Hypotonic Cell(s)	Hypertonic Cell(s)
1. The concentration of dissolved substances (stuff) in the solution is lower than the concentration inside the cell.			
2. When a cell is placed in this solution, water will enter the cell by osmosis resulting in osmotic (turgor) pressure.			
3. The concentration of dissolved substances (stuff) in the solution is the same as the concentration inside the cell.			
4. The concentration of dissolved substances (stuff) in the solution is higher than the concentration inside the cell.			
5. When this solution is injected into the body no cell disruption occurs because no net osmosis occurs.			
6. Putting a plant in this solution will result in water loss and cause the plant to wilt.			

7. In a hypotonic solution, animal cells experience osmosis and a pressure builds up in the cell. What causes the pressure? How is the cell potentially affected by this situation?

8. Why does a plant like celery get limp when placed in a hypertonic solution?

9. A salt water protozoan(a single celled protest) is transferred to a freshwater lake. What might happen to the protozoan?

10. In the figure of the U-tube below, side "A" is pure water and side "B" is a 10% starch solution. Why is the level of side "B" rising? Which side is HYPERTONIC compared to the other?



11. Pure water is always what (Hypertonic, Hypotonic, or Isotonic) compared to any cell? Why?

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Spectrum



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