# Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
  - a. Write the balanced equation for the reaction given above:

+ NaNO<sub>1</sub> → Cu(NO<sub>1</sub>)<sub>2</sub> + ChrCile CuCl<sub>2</sub> + 2 NaNO<sub>3</sub> → Cu(NO<sub>3</sub>)<sub>2</sub> + 2 NaCl

 If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

1 mol CuCl<sub>2</sub> x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 15 g CuClax 134,362g CuCl: 1 mol CuCl: 1 mol NaCl

20 g NaNO<sub>3</sub> x 1 mol NaNO<sub>3</sub> x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84,997g NaNOs 2 mol NaNOs 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride
- d. How many grams of copper (II) nitrate is formed?

 $15g CuCl_2 \times 1 \mod CuCl_2 \times 1 \mod Cu(NO_3)_2 \times 187.467 g Cu(NO_3)_2 = 20.9 g Cu(NO_3)_2$ 134,452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol Cu(NO<sub>3</sub>)<sub>2</sub>

e. How much of the excess reasent is left over in this reaction?

15g CuCl<sub>2</sub> x 1 mol CuCl<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 19 g NaNO<sub>3</sub> 134.452g CuCl<sub>2</sub> 1 mol CuCl<sub>2</sub> 1 mol NaNO<sub>3</sub>

20 g NaNOs - 19 g NaNOs = 1 g NaNOs

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is Percent yield =  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$ the percent yield of this reaction?

formed.

11.3/13.0 x 100% = 86.9%

- 2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are
  - a. Balance the following equation:

Pb(NO<sub>3</sub>)<sub>2 (and</sub> + NaI (and → PbI<sub>2 (a)</sub> + NaNO<sub>3 (and</sub> Pb(NO<sub>3</sub>)<sub>2 tant</sub> + 2 NaI <sub>tant</sub> → PbI<sub>2 tot</sub> + 2 NaNO<sub>3 tant</sub>

 If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO<sub>3</sub>)<sub>2</sub> x 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 12.8 g NaNO<sub>3</sub> 331.214 g Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol Pb(NO<sub>3</sub>)<sub>2</sub> 1 mol NaNO<sub>3</sub>

15.0 g Nal x 1 mol Nal x 2 mol NaNO<sub>3</sub> x 84.997g NaNO<sub>3</sub> = 8.51 g NaNO<sub>3</sub> 149,89g Nat 2 mol NaI 1 mol NaNOs

8.51 g NaNO3 can be made

# **Answers Limiting Reactants And Percent Yield**

**SMITH** 

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